

CIA-RDP86-00513R0008259100

USSR/Chemistry - Anthraquinone
Sulfonation

Dec 48

"Study of an Anthraquinone Series: III, Action of
Sulfonating Agents on the Alpha- and Beta-Mercury
Compounds of Anthraquinone," V. V. Kozlov, Moscow
Order of Lenin Chemico-technol. Inst. imeni D. I.
Mendeleev, 84 pp

"Zhur Obshch. Khim." Vol. XVIII, No 12

Fuming H_2SO_4 reacted with 1-anthraquinonylmercury
sulfate to give mainly 1-anthraquinonylsulfonic
acid. Action of chlorosulfonic acid or fuming
 H_2SO_4 plus $NaCl$ on 1-anthraquinonylmercury sulfate

67/49736

USSR/Chemistry - Anthraquinone (Contd.) Dec 48

gave anthraquinone. Fuming H_2SO_4 reacted with
1,1'- and 2,2'-dianthraquinonyl-mercury to give
mainly the corresponding monosulfonic acids of
anthraquinone.

67/49736

PA 67/49736

KOZLOV, V. V.

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Anthraquinone compounds. XIII. Action of sulfonating agents on 1- and 2-mercury-substituted compounds of anthraquinone. V. V. Kozlov, Zhur. (Miskhet Khim. [J. Gen. Chem.] 18, 2001-2102(1948); cf. C.A. 43, 2199c.—1-Anthraquinonylmercury sulfate (2.5 g.) after 1 hr. at 100° with 8.8 g. 5% oleum gives 39% anthraquinone and 34% K 1-anthraquinonesulfonate, isolated by salting out with KCl; 15% oleum gives 64% of the latter as well as traces of the 1,5- and 1,8-disulfonates; sulfonation with 15% oleum at 120° in the presence of a little NaCl 1.5 hrs. gave 16% anthraquinone and 75% 1-sulfonic acid, as well as 5% 2-sulfonic acid (isolated by pptn. by diazotized benzidine in the form of a salt $[C_{14}H_9(CO)_2SO_3N_2 \cdot C_6H_5]_2$, pptd. quantitatively in acid soln.); sulfonation by $ClSO_3H$ 3 hrs. at 135° gives 78% anthraquinone, with traces of the 2-sulfonyl chloride (sepd. by crysln. from petr. ether), 10% K 1-anthraquinonesulfonate, and 9% 2-sulfonate; excess $ClSO_3H$ gives 39% anthraquinone, 30% 2-anthraquinonesulfonyl chloride, m. 103°, 10% 1-anthraquinonesulfonate, and 9% 2-isomer. Sulfonation of 2.4 g. $[1-C_6H_4(CO)C_6H_4]_2Hg$ by 10 g. 20% oleum 3 hrs. at 135° gave 0.84 g. K 1-anthraquinonesulfonate (from H_2O); the insol. matter on extn. with AcOH gave 0.076 g. anthraquinone and 1.6 g. 1-anthraquinonylmercury sulfate; similar reaction with $ClSO_3H$ gives much $HgCl_2$, a little anthraquinone, and 25% 1-chloroanthraquinone, m. 158°, the residue being 1-anthraquinone-sulfonate (about 55%). Similar sulfonation of 2.4 g. $[2-C_6H_4(CO)C_6H_4]_2Hg$ with 20% oleum gave 1.88 g. mixed Na 2- and 1-anthraquinonesulfonates (contg. 6%, and 94% of the isomers, resp.), 9% anthraquinone, and a small amt. of 2-anthraquinonylmercury sulfate; similar reaction with $ClSO_3H$ gave $HgCl_2$, while the H_2O -insol. solid on extn. with $CHCl_3$ gave about 50% unchanged R_2Hg , obtained in pure state after washing with hot PhOH; the cooled PhOH soln. gave a small amt. of 2-anthraquinonylmercury chloride, infusible prisms, which with $KClO_4$ in AcOH gave 2-chloroanthraquinone, m. 207.5°; the $CHCl_3$ ext. (see above) gave 20% 2-chloroanthraquinone and a small amt. of 2-anthraquinonesulfonyl chloride, m. 102° (from petr. ether). G. M. Kosolapoff

AS 5-51.4 - CHEMICAL LITERATURE CLASSIFICATION

USSR/Chemistry - Anthraquinonedisulfonic Acid Jul 48

Chemistry - Hydrolysis

"Hydrolysis of Anthraquinonedisulfonic Acids," V. V. Kozlov, Moscow Chemicotech Inst imeni D. I. Mendeleev, 4 pp

"Dok Ak Nauk SSSR" Vol IXI, No 2

Experiments with disulfonic acids of anthraquinone and their sodium salts in 80% sulfuric acid and in water in sealed tubes at 260° showed that their conversion resembles that of sulfonic acids of other aromatic compounds, but requires more severe conditions.

11/4979

USSR/Chemistry - Anthraquinonedisulfonic Acid (Contd) Jul 48

Either one or both sulfo groups are substituted by hydrogen, regardless of their position (α , α' or β , β'), and mercury salts promote the process. Concentration of these salts and temperature control extent of hydrolysis. In sealed tubes sulfo groups are substituted by a hydroxyl. Hydrolysis of the sodium salts, especially in water, is 2 - 2.5 times slower. The closer the two sulfo groups are to one of the carbonyl groups, the slower is the hydrolysis of one of them, and the more pronounced is the prevailing direction of the process. Submitted 28 Apr 48.

11/4979

PA 11/4979

KOZLOV, V. V.

KOZLOV, V. V.

Kozlov, V. V. - "Seventy years from the birth of the honorary member of the VKhO (All-Union Chemical Society) imeni D. I. Mendeleev Vladimir Mikhaylovich Rodionov", Soobshch. o nauch. rabotakh chlenov Vsesoyuz. khim. o-va im. Mendeleeva, 1949, Issue 1, p. 1-4, with portrait.

SO: U-4630, 16 Sept. 53, (Letopis 'Zhurnal 'nykh Statey, No. 23, 1949).

P_2O_5 and is very hygroscopic; its di-K, di-Na, di-(NH₄), di-Ag, and Mg salts are sol. in H₂O; Cu (3H₂O), yellow plates made from di-(NH₄) salt: Ca (3H₂O) and i:116 at 18°; Ba (from H₂O or EtOH), sol. i:125 at 21° and i:116 at 18°; Sr (from H₂O or EtOH), scales (from H₂O), plates (from H₂O or H₂O), Ni (5H₂O), scales (from H₂O), plates (from H₂O or H₂O); Pb (H₂O), plates (from H₂O) at 100°; Co at 20° and i:166 at 100°; Zn (H₂O), plates (from H₂O or dil. AcOH), sol. i:650 at 20° and i:275 at 100°; Cu (4H₂O), crystals (from H₂O), sol. i:850 at 20°, and i:108 at 100°; Ba (2H₂O), plates (from H₂O or dil. AcOH), sol. i:100°; Ba (2H₂O), plates (from H₂O or dil. AcOH), crystals (from H₂O), sol. i:145 at 100°; Zn (4H₂O), crystals (from H₂O), sol. i:200 at 20° and i:95 at 100°. Similarly, 1,4-antraquinonedisulfonfyl chloride treated with Na₂S and kept 0.5 hr. at 40°, filtered, cooled to 2°, and acidified with concd. HCl, followed by satn. with dry HCl for 15-30 min. with cooling gave a granular 1,4-disulfonic acid, obtained on vacuum drying over NaOH in 90% yield, in acid, obtained on vacuum drying over NaOH in 90% yield, very sol. in H₂O, EtOH, Me₂CO, somewhat sol. in xylene, and almost insol. in hydrocarbons; most of its salts are sol. in H₂O. The acids have the following m.p.s.: 1,5, 220°, 1,8, 171°, and 1,4, 164°. XV. 1,5-Mercury-containing unsymmetric compounds of anthraquinone. Ibid. 745-51.—1,5-Diaminoanthraquinone (2.38 g.) ground with 2 ml. glycerol was mixed with 3 ml. concd. HCl, refluxed 30 min., cooled to 40-50°, diazotized with 3 ml. 30% NaNO₂, filtered after 2 hrs., and treated with ice water and EtOH, and vacuum-dried represented a 47% yield of HgCl₂ salt of the diazonium deriv., C₁₄H₉N₂Cl₂. This Hg₂ decomp. 330°, sol. in H₂O, EtOH, Me₂CO. This (8.7 g.) suspended in 2.5 g. Cu bronze, let stand overnight, filtered, the ppt. washed with hot EtOH, and the filtrate treated with charcoal and concd. gave 9.8 g. crude

product (m. 136-44°), purified by 40 hrs. extn. with hot-xylene and concn. until it m. 230-41° (it was still impure); extn. for 5 days with cold EtOH gave some 1,5-dichloroanthraquinone, m. 243°, while the residue contained Hg and Cl; treatment of the xylene extn. residue with 30% HNO₃ 12 hrs. and washing with H₂O and hot EtOH gave 27% 1,5-bis(chloromercuri)anthraquinone (I). This (1.0 g.) suspended in 15 ml. KOH in 60 ml. dry MeOH and refluxed 6 hrs. gave a brown-black ppt. (90.4%) of 1,5-CuH₂O₂ (HgOH), insol. in the usual solvents. This (0.5 g.) added to 10 ml. stirred concd. H₂SO₄ at 0°, filtered, and carefully dild. to 3 vols. with H₂O gave 85.5% 1,5-CuH₂O₂(H₂SO₄). I, sol. only in weak oleum. I is recovered unchanged after 1.5-2 hrs. in 5% oleum at 15°, but 10% oleum gives upon treatment with ice water 80% of a mixed salt, apparently an (HgCN)(HgSO₄)H₂ deriv., decomp. 350°; 10% oleum 6 hrs. at 30° gave 94.5% of the di-(HgSO₄)H₂ deriv., which is obtained in 85.9% yield with 15% oleum 6 hrs. below room temp. The product (4 g.) treated with 12 g. 10% oleum 30 min. at 140° and salted out with KCl gave 0.56 g. salt of a disulfonic acid, which yielded 1,5-dichloroanthraquinone, m. 241°, and 2.8 g. Hg deriv. was recovered; 4.8 g. 25% oleum 1.5 hrs. at 140° gave 0.06 g. crude anthraquinone (insol.) and the mother liquor on salting out gave 2 g. di-K anthraquinonedisulfonate, yielding pure 1,5-dichloroanthraquinone (identifying the acid as the 1,5-isomer). XVI. 1,8-Mercury derivative of anthraquinone (unsymmetric). *Ibid.* 803-8; cf. Danilov and Koz'mina, C.A. 43, 6570c.—Di-Na 1,8-

anthraquinonedisulfonate (1.04 g.) in 100 ml. 0.1 N NaOH added to 5 g. HgCl₂ in 20 ml. hot H₂O, boiled 3 hrs., with addn. of 0.1 N NaOH until the soln. is clear initially, gave 2.87 g. 1,8-bis(chloromercuri)anthraquinone, R(HgCN), which (1 g.) suspended in 15 g. KOH in 60 ml. dry MeOH and refluxed 6 hrs. gave 86.16% of the (hydroxymercuri) analog, green-grey powder. The latter (0.5 g.) in 10 ml. concd. H₂SO₄ mixed at 0° and carefully dild. with 3 vols. H₂O gave 88% 1,8-CuH₂O₂(H₂SO₄), an infusible solid; further dildn. of the filtrate gave some anthraquinone; the same sulfate is obtained in 77% yield upon keeping the HgCl deriv. in 10% oleum 6 hrs. at 30-5° or in 18% oleum 6 hrs. in the cold (73% yield). The sulfate (1 g.) treated with 12 g. 10% oleum at 140° 0.5 hr. gave 0.29 g. 1-anthraquinonesulfonic acid (as the chrysoidine salt), and 0.45 g. 1,8-anthraquinonedisulfonic acid (isolated as the di-K salt by salting out), while extn. of the residue with AcOH gave 2 g. unreacted sulfate and 0.3 g. of apparently 1-sulfonanthraquinonylmercury sulfate (assumed to be the 8-isomer), brown, darkening at 325° (from dil. AcOH). Sulfonation with 25% oleum at 140° 1.5 hrs. gave 0.18 g. 1-anthraquinonesulfonic acid, 0.10 g. di-K 1,5-disulfonate, and 1.71 g. di-K 1,8-disulfonate; the insol. residue was free of Hg and gave pure anthraquinone on distn. Hence, the 1,8-(bimercury sulfate) gives a wider variety of products on sulfonation than the 1,5-isomer because of greater thermal lability of the Hg intermediate. G. M. Kosolapoff

PA 65/49T30

KOZLOV, V. V.

USSR/Chemistry - Anthraquinone
Diaz Reaction

Apr 49

"Research on the Anthraquinone Series: IV, The
1, 5-Mercury-Containing Nonasymmetrical Compounds
of Anthraquinone," V. V. Kozlov, D. D. Smolin,
6 1/2 pp

"Zhur Obshch Khim" Vol XIX, No 4

By the A. N. Nesmeyanov diazo method, prepared
1, 5-anthraquinonyl mercury and its corresponding
sulfate derivatives, the latter of which was re-
acted with oleum to form the 1, 5-disulfo acid of
anthraquinone. Submitted 31 Jan 48.

65/49T30

KOZLOV, V. V.

PA 67/49T55

USSR/Chemistry - Anthraquinone
Mercury Compounds

May 49

"Research on the Anthraquinone Series: XVI, 1,8-Mercury-Containing Nonsymmetrical Compounds of Anthraquinone," V. V. Kozlov, D. D. Smolin, 5¹/₂ pp

"Zhur Obshch Khim" Vol XIX, No 5

Synthesizes hydrochloride of 1,8-anthraquinonyl-mercury and its several derivatives, and determines that the principal product formed from the activity of oleum on this compound or the corresponding sulfate is 1,8-disulfo acid of anthraquinone. Submitted 31 Jan 48.

67/49T55

KOZLOV, V.V.

Formation and hydrolysis of anthraquinonesulfonic acids. Voprosy Anilino-
krasochnoy Khim., Trudy VIII Soveshchaniya Khim. i Tekh. (Trans. 8th Aniline
Dye Conf.) '50, 167-80. (MLRA 4:4)
(CA 47 no.21:11172 '53)

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The fastness to light of madder lakes of alizarin and some of its derivatives. V. V. Kozlov and K. M. Salova *J. Applied Chem. U.S.S.R.* 23:369 42 (1950) (Engl. translation); *Zhur. Priklad. Khim.* 23, 317 28 (1950). Madder lakes prepd. from pure, individual anthraquinone dyes in which the substances responsible for lake formation are attached to only one half of the ring were investigated to det. fastness to light. The following represents the diminishing fastness to light: purpurin-3-carboxylic acid > purpurin-3-sulfonic acid > purpura > alizarin > quinizarin > anthragallol > 1,2,3,4-tetrahydroxyanthraquinone. New salts of tetrahydroxyanthraquinone and purpurin-3-carboxylic have been synthesized. M. McMahon

1951

KOZLOV, V. V.; STEPANOV, B. I.

Diethylaniline

Effect of glycerol and saccharose on the reaction rate of p-nitrobenzene diazonium with diethylaniline. Zhur.fiz.khim., 16, No. 6, 1952.

Monthly List of Russian Accessions, Library of Congress, November 1952. Unclassified.

KOZLOV, V. V.

Chemical Societies

Activities of the D. I. Mendeloyev All-Union Chemical Society Vest. AN SSSR 22 No. 1, 1952

Monthly List of Russian Accessions, Library of Congress, June 1952. UNCLASSIFIED.

KOZLOV, V.V.; STEPANOV, B.I.

Influence of pyridine on the rate of coupling of *p*-nitrobenzenediazonium chloride and diethylaniline. Zhur. Fiz. Khim. 28, 592-7 '52. (MLRA 5:6)
(Ca 47 no.20:10494 '53)

1. D.I.Mendeleev Chem.-Technol. Inst., Moscow.

KOZLOV, V.V.; STEPANOV, B.I.

Rate of coupling of *p*-nitrobenzenediazonium chloride with diethylaniline
in the presence of quiholine. Zhur. Fiz. Khim. 26, 701-2 '52. (MLRA 5:8)
(CA 47 no.20:10494 '53)

1. Mendeleev Chem.-Technol. Inst., Moscow.

Organic Chemistry - 10

CP

Effect of glycerol and sucrose on the rate of coupling *p*-nitrobenzenediazonium chloride with diethylaniline. V. V. Kozlov and B. I. Stepanov (D. I. Mendeleev Chem. Technol. Inst., Moscow). *Zh. Fiz. Khim.* 26, No. 8 (1952), 1701. The rate const. of this coupling at pH 4.5 and 15° was 7.30, 10.38, 21.26, and 26.10 $\times 10^{-4}$ l./mol./min. in the presence of 0, 0.204, 0.727, and 1.253 mol. l. glycerol, and 12.08, 23.01, and 53.59 $\times 10^{-4}$ in the presence of 0.0025, 0.100, and 0.500 mol. l. sucrose. Thus, sucrose has the greatest accelerating effect.

J. J. Huker in

KOZLOV, V.V.; STEPANOV, B.I.

Rate of diazotization of *p*-anisidine. Zhur. Fiz. Khim. 27, 3-5 '53.
(CA 47 no.22:11918 '53) (MLRA 6:2)

1. D.I.Mendeleev Chem.-Technol. Inst., Moscow.

KOZLOV, V.V.

In memory of Academician V.M.Rodionov. Ukr.khim.zhur. 20 no.5:
463-469 '54. (MIRA 8:1)

(Rodionov, Vladimir Mikhailovich, 1878-1954)

1917, 1918.

Investigation of the above described documents containing telephone numbers of the
 second time period, October 1944, Bridge St. 77-1010-1011.

20. 11. 1954

Kozlov V.V.

Anthraquinone series. XXI. Hydrolysis of anthraquinone-1-ylmercury chloride. V. V. Kozlov and B. I. Belov. *Zhur. Obshch. Khim.* 25, 410-10 (1955). *J. Gen. Chem. U.S.S.R.* 25, 387-91 (1955) (Engl. translation); cf. *C.A.* 45, 8815g; 44, 7301f. The hydrolysis of anthraquinone-1-ylmercury chloride (I) in H_2SO_4 was studied. The most effective hydrolysis occurs in 75% H_2SO_4 ; the reaction is slower in more concentrated acid and is almost without progress in anhyd. H_2SO_4 . The hydrolysis is retarded by $HgSO_4$ owing to a competing mercuriation reaction. It is considerably affected by temp.: in 75% H_2SO_4 in 1 hr. 35% I is attacked at 100°, nearly 50% at 160°, and over 80% at 190°. The results are shown graphically. $ClSO_3H$ or HCl accelerates the hydrolysis. Diazotization of 1-aminoanthraquinone in

HNO_3 with dry $NaNO_2$ at 0-5° gave solids, of the diazonium salt, which with yellow HgO in HCl yielded anthraquinone-1-diazonium nitrate $HgCl_2$ complex, yellow, decomp. 153°; the corresponding chloride, yellow, decomp. 104°; the sulfate, yellow, decomp. 164°. These double salts couple readily with β -o components. Treatment of any of them with Cu bronze in dry Me_2CO 15 hrs. at 25° gave 41-53.5% 1-chloroanthraquinone and 43-59% I, purified by crystn. from $PhOH$. **XXII. Hydrolysis of anthraquinone-1-ylmercuric sulfate.** *Ibid.* 565-71. Heating anthraquinone-1-ylmercuric sulfate ($RHgSO_4H$) in H_2SO_4 of various concns. was studied. The hydrolysis is most complete in 75-80% H_2SO_4 at 170°. It is hindered by addn. of $HgSO_4$ owing to competing reaction of mercuriation of the resulting anthraquinone. Heating $RHgSO_4H$ in concd. H_2SO_4 yields at 170° anthraquinone and H_2SO_4 , while above 170° there occurs sulfonation of anthraquinone to β -sulfonic acid. The results indicate that $RHgSO_4H$ is an intermediate in the hydrolysis of α -anthraquinonesulfonic acid in the presence of

Hg. XXIII. Hydrolysis of anthraquinone- α -sulfonic acid. V. V. Kozlov and A. A. Egorova. *Zhur. Obshch. Khim.* 25, 800-14; *J. Gen. Chem. U.S.S.R.* 25, 775-9 (1955) (Engl. translation). Anthraquinone- α -sulfonic acid (I) hydrolyzes to anthraquinone in aq. H_2SO_4 at 100-230° under pressure in the absence of Hg . Hydrated Hg salts aid the hydrolysis by forming an intermediate anthraquinone-1-ylmercuric sulfate. At 230° I is hydrolyzed 22-64% in 6 hrs. in 60-85% H_2SO_4 ; the Na salt under such conditions is 81-4% hydrolyzed. Prolongation of hydrolysis from 1 to 18 hrs. and raising the temp. from 190 to 230° in 85% H_2SO_4 lowers the yield of anthraquinone to some 38% as a result of resulfonation to the β -sulfonic acid, which is readily isolated from the hydrolyzates. While 70% H_2SO_4 does not sulfonate anthraquinone, 85-100% H_2SO_4 does react at 210-30°. Addn. of $HgSO_4$ to 85% H_2SO_4 makes possible hydrolysis with 85% H_2SO_4 at 170-50° in an open app., or with 5% H_2SO_4 at 230° in closed app. when 79-88% yields of anthraquinone are formed in 6 hrs.

G. M. Kozlov

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KOZLOV, V.V.; BELOV, B.I.

Investigations in the anthraquinone series. Part 22. Hydrolysis
of α -antraquinonylmercury sulfate. Zhur.of.khim. 25 no.3:565-571
Mr '55. (MLRA 8:7)

1. Institut narodnogo khozyaystva imenii G.V. Plekhanova.
(Hydrolysis) (Mercury organic compounds) (Anthraquinone)

KOZLOV, V.V.; YEGOROVA, A.A.

Study of the anthraquinone series. Part 23. Hydrolysis of anthraquinonesulfonic acid- α . Zhur.ob.khim. 25 no.4:809-814 Ap '55.
(MLRA 8:7)

(Anthraquinonesulfonic acid)

KOZLOV, V. V.

3

Anthraquinone series XXIV. Hydrolysis of anthraquinone- α -sulfonic acid with replacement of the sulfo group by hydroxyl. V. V. Kozlov and A. A. Bessonov. *Zhur. Obshchei Khim.* 29, 661-1000 (1955); cf. C.A. 50, 2490h. Anthraquinone- α -sulfonic acid (I) is hydrolyzed at above 190° by H₂O or dil. H₂SO₄ in a closed system without participation of Hg catalysts yielding α -hydroxyanthraquinone; at higher concn. (85%) of H₂SO₄, anthraquinone also appears. Addn. of Hg salts does not alter the direction of the reaction of hydrolysis, but 5% or more Hg salts favors the formation of anthraquinone rather than of hydroxyanthraquinone. At 230° in H₂O the rate const. is 2.27×10^{-4} in a monomol. reaction; at 190° the const. is 0.2×10^{-4} ; at 210° 1.02×10^{-4} ; at 230° 2.62×10^{-4} ; at 300° 6.21×10^{-4} . Kinetic data are supplied. Also

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in *J. Gen. Chem. U.S.S.R.* 25, 1153-7 (1955) (Engl. translation). XXV. Hydrolysis of anthraquinone-1,8-disulfonic acid. V. V. Kozlov. *Ibid.* 1205-12; cf. C.A. 43, 3400g. Anthraquinone-1,8-disulfonic acid (I) can be hydrolyzed in H₂O and H₂SO₄ at 230° or higher in a closed system without the presence of Hg salts. Either one or both sulfonic acid groups are replaced by H or HO, all the combinations being found among the products. Hg salts aid the formation of products in which II replaces the SO₃H group. The most complete hydrolysis occurs in 0.25% H₂SO₄ when in 6 hrs. at 200° 100% 1-hydroxyanthraquinone-8-sulfonic acid (II), m. 256° (from HCl), is formed; this forms a sparingly sol. Na salt. The amt. of unreacted I can be detd. by treatment of the mixt. with KClO₄; the I remains unoxidized. II is not chlorinated by KClO₄. The acidic groups in I are thus not equivalent. The kinetic results of hydrolysis with and without the presence of HgSO₄ are given in tables and graphically. G. M. Korshakoff.

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KOZLOV, V.V.

Investigation of the anthraquinone series. Part 25. Hydrolysis of
1,8 disulfoacids of anthraquinone. Zhur.ob.khim.25 no.6:1206-1212
Je'55. (MIRA 8:12)

1. Institut narodnogo khozyaystva SSSR.
(Anthraquinonesulfonic acid)

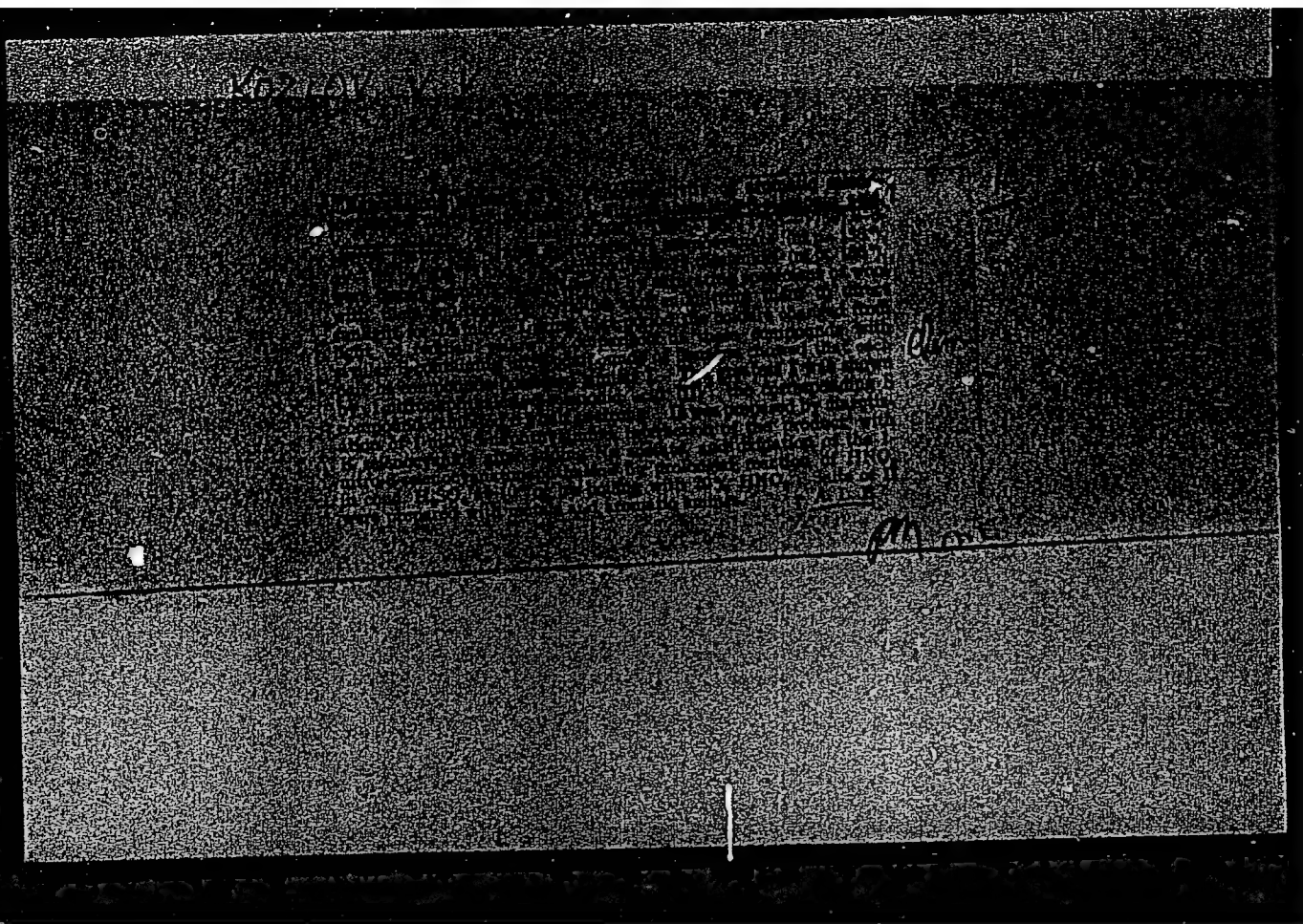
SPITSYN, V.I.; KAPUSTINSKIY, A.F.; KOZLOV, V.V., doktor khimicheskikh nauk.

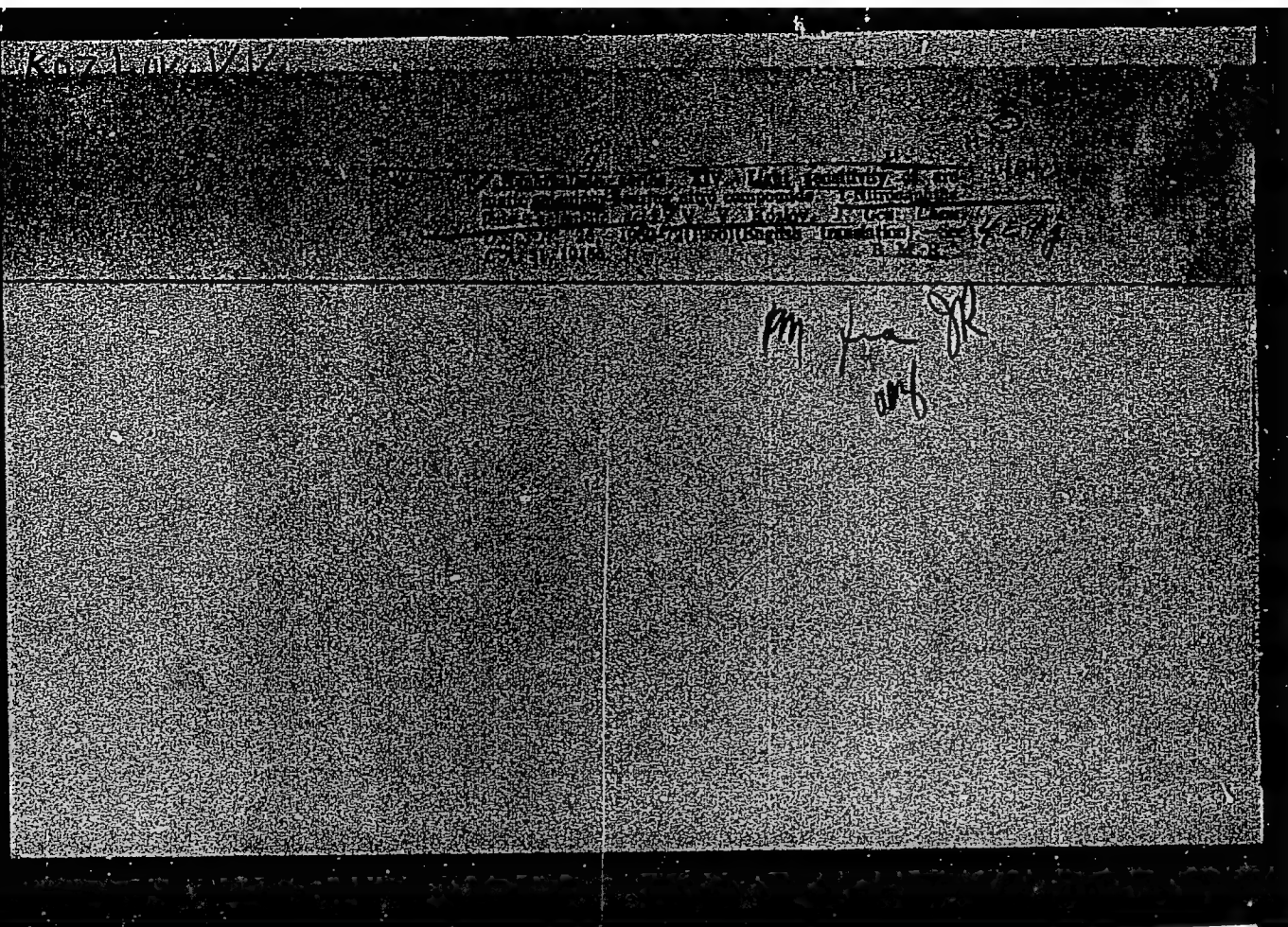
Hungarian chemical congress. Izv.AN SSSR.Otd.khim.nauk no.5:635-640
My '56. (MIRA 9:9)

1.Chlen-korrespondent AN SSSR (for Spitsyn, Kapustinskiy).
(Hungary--Chemistry--Congresses)

KOZLOV, V.V.

Vasilii, Vasil'evich Sharvin. Trudy Inst.ist.est.i tekhn. 12:285-311
'56. (MLRA 9:12)
(Sharvin, Vasilii Vasil'evich, 1870-1930)





Kozlov, V. V.

USSR/Organic Chemistry. Synthetic Organic Chemistry.

E-2

Abs Jour: Ref Zhur-Khimiya, No 6, 1957, 19164.

Author : Kozlov V. V.

Inst :

Title : Studies in the Anthrachinon Series. XXVI. Hydrolysis of 1,5-disulfonic acid of Anthrachinone.

Orig Pub: Zh. obshch. khimiyi, 1956, 26, No 9, 2617-2622

Abstract: By the hydrolysis of 1,5-disulfonic acid of anthrachinon (I) and its Na- and K-salts in sealed tubes (0.0025 mole substance, 25 cc H₂SO₄ or water) all theoretically possible products of the hydrolysis of I, including the products of the substitution of SO₃H-group for H were isolated. The amount of I hydrolyzed in 6 hours, by increasing the temperature from 230 to 260°, in 80% H₂SO₄ is increased 2 times, in 20%---3 times, in a 0.1%--- and 9% to 98.2%. Hydrolysis of I in water (260°, 6 hours)

Card : 1/2

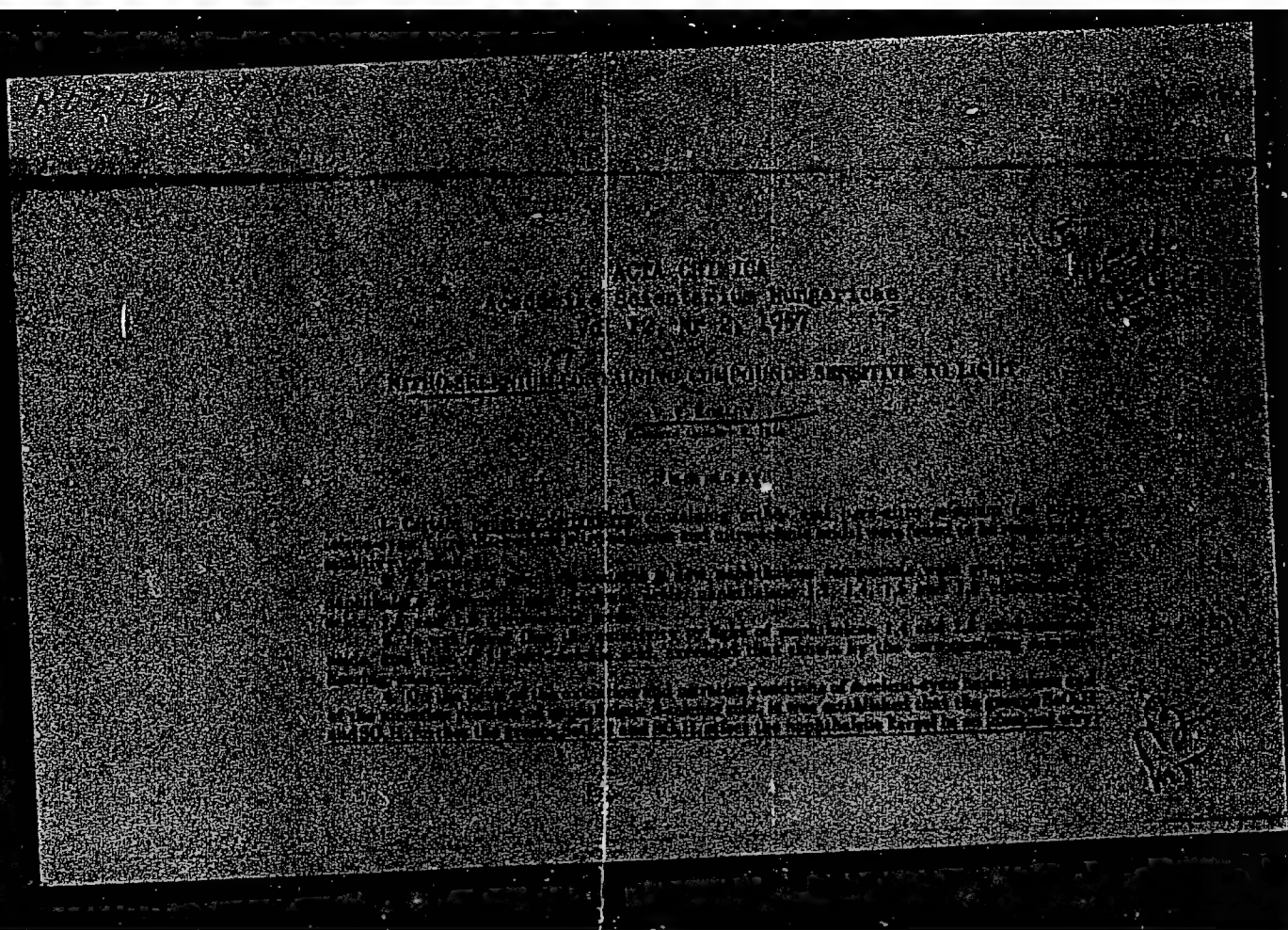
USSR/Organic Chemistry. Synthetic Organic Chemistry.

E-2

Abs Jour: Ref Zhur-Khimiya, No 6, 1957, 19164

proceeds to 97.2%. Thereby are formed Na-salt of 1-hydroxysulfonic acid-5 (II) 21%; -sulfonic acid 36.4%; hydroxyanthrachinon 10.8%; anthrachinone (III) 28.6%. In H₂SO₄, as well as in water, the sulfo group is substituted in a higher degree by H than by OH. The Na and K-salts of I do not hydrolyze in water but in H₂SO₄ they hydrolyze with greater difficulty, than does the free I. At the hydrolysis of Na-salt of I in a 70% acid (260°, 6 hours), 87% II and 12% III are formed. The addition of H₂SO₄ lowers the temperature, at which the hydrolysis occurs and facilitates the formation of III. Report XXV see RZhKhim, 1956, 22435.

Card : 2/2



2-Naphthylamine (XV). Rearrangement of aromatic amines and formation of the dyestuffs of 2-naphthylamine with the 2-naphthylamine acid. V. V. Kozlov and G. G. Kozlov, *Izv. Akad. Nauk SSSR Khim. Biol. Nauk*, 1964, 1360. (Chem. Abstr., 1965, 59:13604c.)
A mixture of 2-naphthylamine (1 mole) and 10% NaOH solution (2.5 moles) results in 10% transfer of the amino group to 2-naphthylamine acid which is stable to acids and alkalis. At higher temps. the rearrangement proceeds to the salt of 2-bromo-2-naphthylamine-6-sulfonic acid (m.p. 200°C) which is precipitated in the reaction mix. (acids) and this compound can be obtained in the reaction in a facile way. The nature of the reaction mixture of the reaction is the subject of study. The reaction mixture was analyzed as follows: 1) with HCl and CaH₂ with acid; with CaH₂ and the mixture was treated with 0.1N NaOH. The residue after acid was treated with 10% HCl to hydrolyze the ester group and the resulting amine was titrated as above. As the part of the residue was titrated with 0.1N NaOH after addition in five the sum of 2-naphthylamine and 2-naphthylamine. The former was acid by treatment with HClO which yields 2-bromo-2-naphthylamine in 80%. 2-Naphthylamine-6-sulfonic acid was isolated as the Na salt which gives a 1:10 diazonium salt; it was quantitatively oxidized to 2,6-dichloro-2,6-diamine. XVI. Transformation of the salt of 2-naphthyl-6-sulfonic acid into salt of 2-naphthyl-6-sulfonic acid. V. V. Kozlov and G. G. Kozlov, *Izv. Akad. Nauk SSSR Khim. Biol. Nauk*, 1964, 1360. (Chem. Abstr., 1965, 59:13604c.)
2-Naphthylamine (1 mole) and 10% NaOH solution (2.5 moles) in 100% (anhydrous) sulfuric acid in 215°C. heated in a mix. with acid in 210-20° yields 2-naphthyl-6-sulfonic acid. The reaction apparently proceeds through formation of a sulfonate and probably occurs stepwise. 2-Naphthyl-6-sulfonic acid crystals with 2.5% HCl. The acid in 215°C. heated in 210-20°C. in 185°. The reaction shows the importance of the amino group to substitute in the aromatic ring in the nucleus. G. M. K.

Dieter HALL/DEW

KOZLOV, V.V.

KOZLOV, V.V.; YAKOBSON, G.G.

Research on naphthalene series. Part 16: Conversion of 2-naphthol-6-sulfonic acid salt into 2-naphthol-7-sulfonic acid salts. Zhur.
ob.khim. 27 no.5:1156-1160 My '57. (MIRA 10:8)
(Naphtholsulfonic acid)

Kozlov, V. V.

79-11-22/56

AUTHOR: Kozlov, V. V.

TITLE: Investigations in the Naphthalene Series (Issledovaniya v naftalinovom ryadu).
XVII. On the Sensitivity to Light of the Aromatic Selenium-Nitro-Compounds: 1-Nitronaphthalene-5-, 1-nitronaphthalene-2- and 1-nitronaphthalene-4-Selenic Acids (XVII. O svetochuvstvitel'nosti aromaticeskikh selensoderzhashchikh nitrosoyedineniy. 1 - nitronaftalin- 5 -, 1 nitronaftalin- 2 i 1 nitronaftalin - 4 - selenistyie kisloty).

PERIODICAL: Zhurnal Obshchey Khimii, 1957, Vol. 27, Nr 11, pp. 3011-3015 (USSR)

ABSTRACT: In addition to his earlier investigation the author synthesized other isomeric nitroselenic acids of naphthalene. He obtained them according to the general scheme by conversion of the diazotized isomeric nitronaphthalenes (see chemical formula). The diazotization is best performed in an acetic solution with dry nitrite, and the subsequent oxidation of the selenium compound is best performed by boiling with 20% nitric acid. The aqueous solutions of the

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79-11-22/56

Investigations in the Naphthalene Series.
XVII. On the Sensitivity to Light of the Aromatic
Selenium-Nitro-Compounds: 1-Nitronaphthalene-5-,
1-nitronaphthalene-2- and 1-nitronaphthalene-4-Selenic Acids

isomeric nitroselenic acids of naphthalene and the woolen tissues soaked with them behave differently toward the sunlight. The quantities of the sensitivity to light of the isomeric nitroselenic acids of naphthalene, in comparison with their sulfur-containing analogues, are to be seen from the table. Thus isomeric selenic acids of nitronaphthalene (1,2; 1,4; 1,5) and their salts were obtained. It was found that the 1,5-nitroselenic acid of naphthalene is sensitive to light, and that the 1,2-acid is slightly sensitive. The 1,4-nitroselenic acid of naphthalene shows a marked sensitivity as compared to its sulfur-containing sister 1-nitro-4-sulfinic acid of naphthalene. There are 1 table, and 3 references, all of which are Slavic.

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Investigations in the Naphthalene Series.

79-11-22/56

XVII. On the Sensitivity to Light of the Aromatic Selenium-Nitro-Compounds: 1-Nitronaphthalene-5-, 1-nitronaphthalene-2- and 1-nitronaphthalene-4-Selenic Acids

ASSOCIATION: Institute of Economics imeni G. V. Plekhanov
(Institut narodnogo khozyaystva im. G. V. Plekhanova).

SUBMITTED: November 23, 1956

AVAILABLE: Library of Congress

1. Aromatic selenium - Nitro-compounds - Light sensitivity
2. Light - Chemical effects

Card 3/3

KOZLOV, Vladimir Veniaminovich, prof.; VOL'FKOVICH, S.I., akademik, red.;
BANKVITSER, A.L., red. izd-va; POLYAKOVA, T.V., tekhn.red.

[History of chemical societies of the U.S.S.R.] Ocherki istorii
khimicheskikh obshchestv SSSR. Pod red. S.I. Vol'fkovicha.
Moskva, Izd-vo Akad. nauk SSSR, 1958. 609 p. (MIRA 12:1)
(Chemical societies)

KOZLOV, V.V., prof.

~~Eighth Mendeleev Congress. Khim.nauk i prom. 3 no.5:546-550 '58.~~
(MIRA 11:11)

1. Glavnyy uchenyy sekretar' Orgkomiteta VIII Mendeleevskogo
s"yezda.
(Chemistry--Congresses)

APPROVED FOR RELEASE: Monday, July 31, 2000

CIA-RDP86-00513R00082

100

5(2,3)

AUTHORS:

Khrustaleva, V. N., Kozlov, V. V.

TITLE:

Colorimetric Determination of Soluble Carbohydrates by Means
of Heteropolyacids (Kolorimetriceskoye opredeleniye
rastvorimyykh uglevodov s pomoshch'yu geteropolikislot)

PERIODICAL:

Izvestiya vysshikh uchebnykh zavedeniy. Khimiya i khimicheskaya
tekhnologiya, 1958, Nr 6, pp 47-53 (USSR)

ABSTRACT:

According to the findings made by the authors, phospho-
molybdic acid can be employed, not only for the purposes
already known (Refs 1-3), but also, in a sulphuric acid
solution on short heating with carbohydrates, for the deter-
mination mentioned in the title, due to the fact that it
develops an intensively blue coloring (Ref 4). The paper
under consideration deals with the application of this re-
action to the purpose mentioned in the title. A recipe is
given. Although the interaction of the above acid takes place
even in the absence of H_2SO_4 in the reaction medium, the
latter increases the coloring intensities of the solutions.
In this connection, heteropolyacid, via several processes,
is brought to a low molybdenum valency (Refs 5-10). The limits
of the quantitative glucose determination are between 3 and

SOV/153-58-6-9/22

Colorimetric Determination of Soluble Carbohydrates by Means of Heteropolyacids

42 g/liter (0.016-0.23 mol/liter). The sensitivity of this reaction with regard to fructose (bluish-green coloring) is between 0.3 and 15 g/l (0.0016 - 0.083 mol/l). In glucose-fructose solutions, the same coloring intensity emerges on a concentration ratio of 10:1. The results are listed in figure 1. On the heating with the acid concerned (in the absence of H_2SO_4) of disaccharides, the coloring appears only in saccharose. H_2SO_4 enhances this coloring. Maltose and lactose yield a coloring only in the presence of H_2SO_4 (Fig 1). The substitution of sulphuric acid by phosphoric acid⁴ enhances the coloring in the glucose-fructose mixture also in the presence of other carbohydrates. The coloring becomes purer, and the reaction reasonably sensitive. Table 1 shows the dependence on the carbohydrate concentrations of the coloring intensity of the solution. From this it can be seen that an even increase of the optical density as a function of carbohydrate concentration is characteristic of the sugar solutions with a keto group, in particular of fructose and saccharose. In this experiment there were noticed differences in the reducing power of the carbohydrates. The sugars belonging to the aldoses or containing an aldo group show but slight reduc-

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Colorimetric Determination of Soluble Carbohydrates by Means of Heteropolyacids

tion under the given conditions. The concentration increase of such carbohydrates but little increases the optical densities. Figure 2 shows the data for keto sugars from table 2 entered into a coordinate system. Here, with regard to fructose, raffinose, and to lower concentrations of saccharose, all points are arranged on a straight line. The curves concerned pass through the coordinate origin. Thus the solutions colored due to the reaction in question obey Lambert-Buger-Beer's law. The accuracy of the results can be increased by the use of an improved apparatus. Table 2 shows the determination of the carbohydrates in their mixtures. Table 3 presents determinations of soluble carbohydrates in aqueous extracts from berries, fruit and vegetables. There are 2 figures, 3 tables, and 12 references, 6 of which are Soviet.

ASSOCIATION: Kafedra organicheskoy khimii; Moskovskiy institut narodnogo khozyaystva imeni G. V. Plekhanova (Chair of Organic Chemistry; Moscow Institute of National Economy imeni G. V. Plekhanov)

SUBMITTED: December 9, 1957

Card 3/3

15(7)

AUTHORS:

~~Kozlov, V. V.~~, Solntseva, R. R.

SOV/64-58-7-8/18

TITLE:

On the Hygroscopicity of Acid Azo Dyes
(O gigroskopichnosti kislotnykh azokrasiteley)

PERIODICAL:

Khimicheskaya promyshlennost', 1958, Nr 7, pp 416-420 (USSR)

ABSTRACT:

The dependence of the moisture absorption of dyes as well as their salts on the relative air-moisture in the case of constant temperature were investigated, and the function of this moisture absorption versus time was determined. The maximum moisture content observed during the experiment was regarded as the equilibrium moisture of the dye. The determination of the hygroscopicity was carried out according to the static (or exsiccator) method (Ref 2). The quantity of the equilibrium moisture is a function of the relative air-moisture and the nature of the dye. The sorption isothermal lines have an S-shape similar to those of capillary porous bodies (Ref 3). The moisture of the dye increases at constant temperature with the increase of the relative air-humidity. This increase is especially high at an air-humidity of more than 70%. At the same relative air-humidity (e. g. 70%) the dyes acid red 2 C (15.7%),

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On the Hygroscopicity of Acid Azo Dyes

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acid blood-red (12.9%), acid light red (11.3%) have the highest equilibrium moisture, whereas acid orange (8.6%) has the lowest value. A comparison demonstrates that the order of hygroscopicity of the dyes is almost the same as the order of the solubility. However, no theoretical (Ref 5) dependence on the structure of the dye could be found. The experimental results on the moisture absorption of sodium, potassium, ammonium and calcium salts of the dyes show that (as in the dyes themselves) the moisture for the most part is absorbed within the first days. Fillers increase the hygroscopicity of the dyes. There are 4 figures, 4 tables, and 6 Soviet references.

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KOZLOV, V.V.

On the eve of the Eight Mendeleev Congress. Ukr.khim.zhur. 24 no.5:
690-594 ' 58. (MIRA 12:1)
(Chemistry--Congresses)

DOI / 74-27-6-6/6

AUTHORS: Khrustaleva, V. N., Kozlov, V. V. (Moscow)

TITLE: Color Reactions on Soluble Carbohydrates and Their Utilization in Colorimetry (Tsvetnyye reaktsii na rastvorimyye uglevody i ispol'zovaniye ikh dlya kolorimetrii)

PERIODICAL: Uspekhi khimii, 1958, Vol 27, Nr 6, pp 752-784 (USSR)

ABSTRACT: The author first stress the great importance of color reaction in colorimetric analysis. In this connection an important part is played by carbohydrates because they absorb no light in the visible part of the spectrum. In connection with this reaction, which develops according to the scheme: carbohydrate + oxidizing agent \rightarrow oxidation products of carbohydrate + reduced reagent (colored substance in solution), also indirect reaction is mentioned; it is further pointed out that hitherto comparatively few papers have been published on the colorimetry of soluble carbohydrates. There follows an enumeration of methods, based upon the reaction of carbohydrates with various inorganic compounds. The following problems are discussed in this chapter: 1) The

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Color Reactions on Soluble Carbohydrates and Their Utilization in Colorimetry

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reaction of the reduction of alkaline copper oxide solutions. 2) The reaction of the reduction of acid copper oxide solutions. 3) Reactions with potassium ferrocyanide. 4) Reaction with chromium salts. 5) Reaction with ammonium molybdate. 6) Reaction with heteropoly acids. 7) Reaction with sulphuric acid. 8) Reaction with selenic acid. There follow the reactions with cobalt nitrate sodium nitroprusside, hydrogen bromide, iron chloride, and caustic alkalis. In the second chapter the authors give a survey of the methods based upon the carbohydrate reactions with various organic compounds, viz. reaction with picric acid, dinitrophenol, dinitrobenzene, dinitrobenzoic acid, dinitrosalicylic acid, anthraquinone - nitrosulfonic acid, phenol, thymol, guaicol, resorcinol, resorcin-disulfoacid, orcin, pyrocatechin, floroglucina, α -naphthol, naphthol-sulfoacid, naphtho-resorcinol, chromotropic acid, dinitro-naphthol-sulfoacid, anthrone (or anthranone), methylamine, ethylenediamine, aniline, anisidine, benzidine, diphenylamine, dibenzo-carbazole, thioglycol acid, urea (carbamide), diazourazyl as well as with triphenyltetrasol. In conclusion the authors say that the methods of the colorimetric definition of certain (special) carbohydrates require further elaboration (in their mixtures). There are 283 ref-

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SOV/4-27-6-6/6
Color Functions on Soluble Carbohydrates and Their Utilization in Colori-
metry

exceeds, 25 of which are Soviet.

1. Carbohydrates--Spectrum
2. Colorimetry--Applications
3. Light--Absorption
4. Carbohydrates--Test results

Card 3/3

AUTHORS: Kozlov, V. V., Veselovskaya, I. K. SOV/79-28-12-39/41

TITLE: Investigations of the Naphthalene Series (Issledovaniya v naftalinovom ryadu) XVIII. On the Reaction Kinetics of β -Naphthol With Aqueous Ammonia and Ammonium Sulfite Solution (XVIII. O kinetike vzaimodeystviya β -naftola s vodnymi rastvorami ammiaka i sul'fita ammoniya)

PERIODICAL: Zhurnal obshchey khimii, 1958, Vol 28, Nr 12, pp 3333 - 3338 (USSR)

ABSTRACT: The amination reaction of naphthols with aqueous solutions of ammonia in the presence of salts of sulfurous acid is of practical importance to the synthesis of naphthylamines. The chemism of such a reaction was suggested by N. N. Vorozhtsov (Ref 1) and further explained by the authors according to scheme(1). Naphthol can according to this scheme, enter reaction with sulfites in the hitherto not isolated keto as well as in the enol form producing an affiliation product corresponding to the keto form. The kinetic character of this reaction process has not been described either. The results obtained in the reaction of β -naphthol (melting point 122°) with aqueous ammonia in the presence of ammonium sulfite

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Investigations of the Naphthalene Series. XVIII. On the SOV/79-28-12-39/41
Reaction Kinetics of β -Naphthol With Aqueous Ammonia and Ammonium Sulfite
Solution

are given below in detailed equations. The amination took place in a rotating steel autoclave at 150-180°. The rate of the whole amination process is determined by the first phase, i.e. by the rate of the reaction of β -naphthol with bisulfite (NH_4HSO_3 or NaHSO_3). The affiliation product

of scheme (1) has a high mobility and therefore keeps reacting with ammonia under the formation of naphthyl. amine. The transformation velocity of β -naphthol depends on time and temperature; it is directly proportional to the product of the concentrations of naphthol and bisulfite, and is independent of the concentration of ammonia. The amination velocity can also be caused to depend on the concentration of the decreasing ammonium sulfite. In that case the reaction rate is directly proportional to the product of the concentrations of β -naphthol and sulfite, and is reversely proportional to the concentration of ammonia. Tables, figures and equations illustrate the role played by the concentration of ammonia and β -naphthol, the amount of ammonium sulfite,

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Investigations of the Naphthalene Series. XVIII. On the SOV/79-28-12-39/41
Reaction Kinetics of β -Naphthol With Aqueous Ammonia and Ammonium Sulfite
Solution

time and temperature in the amination of β -naphthol. There
are 2 figures, 5 tables, and 5 references, 3 of which are
Soviet.

ASSOCIATION: Moskovskiy institut narodnogo khozyaystva (Moscow Institute
of National Economy)

SUBMITTED: August 22, 1957

Card 3/3

5(3)

SOV/153-2-3-12/29

AUTHORS: Kozlov, V. V., Belov, B. I.

TITLE: Characteristic Features of Diazotization of Aminoanthraquinones

PERIODICAL: Izvestiya vysshikh uchebnykh zavedeniy. Khimiya i khimicheskaya tekhnologiya, 1959, Vol 2, Nr 3, pp 374-380 (USSR)

ABSTRACT: Various papers with remarks on the mentioned peculiarities are quoted at the beginning. In this connection Lukin (Ref 25) is cited. The authors made comparative investigations on the behavior of isomeric mono- and diaminoanthraquinones in the diazotization in various media. Diazotization is possible in hydrochloric and sulphuric solution; however, different conditions are necessary for different aminoanthraquinones. The addition of potassium bromide produces favorable effects in the diazotization of all aminoanthraquinones. The rate of diazotization of the compounds investigated in hydrochloric and sulphuric solutions decreases in the following series: 1,5-diaminoanthraquinone > 1,8-diaminoanthraquinone > α -aminoanthraquinone > β -aminoanthraquinone > 1,4-diaminoanthraquinone. The diazotization of aminoanthraquinones is possible also in acetic solutions with sodium being used in dry state, without addition of a mineral

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Characteristic Features of Diazotization of
Aminoanthraquinones

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acid. The diazotization takes place very rapidly in glacial acetic acid, somewhat more slowly in 80 % acetic acid. A temperature increase accelerates the diazotization in acetic acid solution. In different aminoanthraquinones the completeness of the reaction depends to a different degree on the duration. The yields of the diazotization of five aminoanthraquinones in acetic acid at different concentrations are listed in table 1. The authors also investigated the stability of the formed diazo compounds. The results are shown in table 2. The following series is obtained for the stability of the diazo compounds of the aminoanthraquinones investigated:
 β - > 1,5- > α - > 1,8- > 1,4. Table 3 shows the decomposition of solutions of diazo- α -aminoanthraquinone in the case of different duration of electric illumination with 500 watt. Table 4 gives a survey on the decomposition of the acetic salts of the diazotized aminoanthraquinones in the case of electrical illumination with 500 w. Table 5 shows characteristics of some double salts of the diazotized aminoanthraquinones with sublimate. The optimum diazotization conditions of α - and β -aminoanthraquinone and of 1,4-, 1,5-, and 1,8-diaminoanthraquinone are

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Characteristic Features of Diazotization
of Aminoanthraquinones

SOV/153-2-3-12/29

mentioned in this paper. Yu. V. Lyande and Ye. K. Malinina participated in the investigations. There are 5 tables and 35 references, 10 of which are Soviet.

ASSOCIATION: Moskovskiy institut narodnogo khozyaystva imeni G. V. Plekhanova - Kafedra organicheskoy khimii (Moscow Institute of Political Economy imeni G. V. Plekhanov - Chair of Organic Chemistry

SUBMITTED: February 11, 1958

Card 3/3

5 (3)

AUTHORS:

Kozlov, V. V., Davydov, A. A.

SOV/32-25-8-9/44

TITLE:

On the Analytic Determination of the Sulfo Acids of the Anthraquinone

PERIODICAL:

Zavodskaya laboratoriya, 1959, Vol 25, Nr 8, pp 926 - 928 (USSR)

ABSTRACT:

It was established that the chlorination reaction needed at the preparation of chloro anthraquinone (I) and at the determination of the sulfo acids of the anthraquinone (SA) (which is based on the same principle), can be simplified. The simplification is achieved by chlorination at the optimum acid concentration which results in the instantaneous mixing of the reagent at the beginning of the process instead of a successive addition of the chlorate. A considerable acceleration of the reaction results both for the α - and β -sulfo-substituted anthraquinones. The article contains the process of analysis of the α -(SA) according to the chlorate method. The chlorination of (SA) in a hydrochloric medium with simultaneous formation of chlorine (Refs 6-7) can also be effected with other oxidation agents ($K_2Cr_2O_7$, CrO_3 , $KMnO_4$, etc). For

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On the Analytic Determination of the Sulfo Acids
of the Anthraquinone

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the analysis of α -(SA) the following reagents are recommended:
1.30 - 1.31 g of the potassium salt of the α -(SA), 100 ml of
HCl (specific weight 1.19), 75 ml of 4n H_2SO_4 , 2.0 g of
 $K_2Cr_2O_7$ in 25 ml of water. The chlorination is rapid and uni-
form. Yield of α -(I) is 98.5 - 99.0% (Table of the reproduci-
bility of the analysis). At the analytical chlorination of
other (SA) (β -, 1.5-, and 1.8-) according to the last-mention-
ed method, the time of adding the $K_2Cr_2O_7$ -solution to the mix-
ture must be prolonged to 2 hours and the mixture has to be
boiled for from 2.5 to 3 hours. In the presence of sulfates
of alkali- and alkaline earth metals weighing more than 25%
of the weight of the (SA) lower results are obtained. In this
case the analysis must be conducted with the use of a mixture
of hydrochloric acid - sulfuric acid. There are 1 table and
7 references, 5 of which are Soviet.

ASSOCIATION: Institut narodnogo khozyaystva im. G. V. Plekhanova (National
Economy Institute imeni G. V. Plekhanov)

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5(0)
 AUTHORS: Kozlov, V. V., Vol'fkovich, S. I. SOV/74-28-12-1/25

TITLE: 8th Mendeleev Congress on General and Applied Chemistry (From March 16 to 23, 1959, in Moscow)

PERIODICAL: Uspekhi khimii, 1959, Vol 28, Nr 12, pp 1399-1403 (USSR)

ABSTRACT: This is a report on the Congress quoted in the title. In the preface, a short chronological survey of the preceding seven Congresses is given. The Congress of Chemists, bearing in honor of D. I. Mendeleev his name, were started by the Russkoye fiziko-khimicheskoye obshchestvo (Russian Physico-chemical Society). On the 6th Congress held in Khar'kov in 1932, the Vsesoyuznoye khimicheskoye obshchestvo im. D. I. Mendeleeva (All-Union Chemical Society imeni D. I. Mendeleev) was established. A. N. Bakh was elected chairman. The 8th Mendeleev Congress was convened by the All-Union Chemical Society imeni D. I. Mendeleev, the Academy of Sciences of the USSR, the Ministerstvo khimicheskoy promyshlennosti SSSR (Ministry of Chemical Industry of the USSR) in cooperation with the Gosudarstvennyy Komitet Soveta Ministrov SSSR po khimii (State Committee of the Council of Ministers of the USSR for Chemistry) ✓

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and the Ministerstvo vysshego obrazovaniya SSSR (Ministry of Higher Education of the USSR) (Ref 2). The Conference was organized under the auspices of the Organization Committee directly subordinated to the Prezidium i Otdeleniye khimicheskikh nauk AN SSSR (Presidium and Department of Chemical Sciences of the AS USSR). Members of the Organization Committee were: Academician A. N. Nesmeyanov (chairman); Corresponding Member AS USSR N. M. Zhavoronkov, Doctor of Technical Sciences I. P. Losev, Academician S. I. Vol'fkovich (all vice-chairmen); Doctor of Chemical Sciences V. V. Kozlov (Scientific Chief Secretary); Academician A. Ye. Arbuzov; Academician B. A. Arbuzov; I. V. Belov of the VSNITO; Academician AS Ukrainskaya SSR A. I. Brodskiy; Academician AS Ukrainskaya SSR P. P. Budnikov; Academician A. P. Vinogradov; Professor G. D. Vovchenko; Corresponding Member AS USSR N. N. Vorozhtsov; Doctor of Chemical Sciences S. V. Gorbachev; Corresponding Member AS USSR S. N. Danilov; Academician M. M. Dubinin; Minister of High Education USSR V. P. Yelyutin; V. A. Ivanov of TsK profsoyuza rabochikh khimicheskoy promyshlennosti (Central Committee of the Labor

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Union of Workers of the Chemical Industry); Professor V. M. Kabadze; Doctor of Technical Sciences V. S. Kiselev; Docent V. P. Komarov; Academician V. N. Kondrat'yev; Academician AS Latvinskaya SSR L. K. Lepin'; Academician AS Azerbaydzhanskaya SSR Yu. G. Mamedaliyev; Doctor of Chemical Sciences K. P. Mishchenko; Corresponding Member AS USSR B. V. Nekrasov; D. P. Novikov of the State Committee of the Council of Ministers of the USSR for Chemistry; Doctor of Technical Sciences A. N. Planovskiy; Deputy Minister of Higher Education USSR M. A. Prokof'yev; Corresponding Member AS USSR O. N. Reutov; Doctor of Technical Sciences Z. A. Rogovin; A. Ya. Ryabenko of the State Planning Office of the USSR; Academician N. N. Semenov; Academician A. N. Terenin; Vice-chairman of the State Committee of the Council of Ministers for Chemistry S. M. Tikhomirov; Academician A. V. Topchiyev; Doctor of Chemical Sciences K. V. Topchiyeva; Member of the Board of the Ministry of Higher Education USSR N. S. Torocheshnikov; Vice-chairman of the State Committee of the Council of Ministers of the USSR for Chemistry G. V. Uvarov; Chairman of the State

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Committee of the Council of Ministers of the USSR for Chemistry V. S. Fedorov; Doctor of Chemical Sciences N. A. Figurovskiy; Academician A. N. Frumkin; Academician I. I. Chernyayev; Academician M. M. Shemyakin; Academician AS Uzbekskaya SSR S. Yu. Yunusov. About 500 persons took an active part in the preparation of the Congress. The 8th Mendeleev Congress was opened on March 16, 1959, in the hall of the Moskovskiy Gosudarstvennyy Universitet imeni M. V. Lomonosova (Moscow State University imeni M. V. Lomonosov) by the chairman of the Organization Committee, Academician A. N. Nesmeyanov. He stated, among other things, that the Academy of Sciences had considerably increased since the last Congress was held. At present, 12 chemical institutes having a broader volume alone than the whole Academy of 1934, belong to it. Chemical science comprises now, in addition, several dozens of industrial scientific research institutes, some hundreds of plant laboratories, and about 80 chemical colleges and departments. Famous scientists of the Soviet Union took an active part in all Mendeleev Congresses. Among these: N. N. Beketov, N. A. Umov, V. I. Vernadskiy, D. P. Kononov, A. Ye. Favorskiy, N. S. Kurnakov, A. Ye. Fersman, ✓

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SOV/74-28-12-1/25

8th Mendeleev Congress on General and Applied Chemistry (From March 16 to 23, 1959, in Moscow)

N. D. Zelinskiy, L. A. Chugayev, V. Ye. Tishchenko, D. N. Pryanishnikov, A. N. Bakh, P. P. Lazarev, V. G. Khlopin, A. A. Baykov, S. I. Vavilov, N. A. Morozov, N. A. Shilov, V. A. Kistyakovskiy, and others. It was stated by the speaker that a delegation of scientists and chemical engineers, members of which were: A. N. Bakh, N. S. Kurnakov, N. D. Zelinskiy, E. V. Britske, A. Ye. Poray-Koshits, V. Ya. Kurbatov, and others, had offered to cooperate with the government on March 14, 1928. For this reason, the Komitet po khimizatsii narodnogo khozyaystva SSSR (Committee for the Realization of Progress in Chemistry in the National Economy of the USSR) was established. After the inaugural discourse was finished, the Prezidium Tsentral'nogo Komiteta Kommunisticheskoy partii Sovetskogo Soyuza (Presidium of the Central Committee of the Communist Party of the Soviet Union) was, on suggestion of the chairman of the Organization Committee Professor I. P. Losev, elected Honorary Presidium of the Congress. There are 1 table and 59 references, 58 of which are Soviet. ✓

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5(9)
 AUTHORS: Kozlov, V. V., Vol'fkovich, S. I. SOV/74-28-12-2/25
 TITLE: 8th Mendeleev Congress. Plenary Meetings
 PERIODICAL: Uspekhi khimii, 1959, Vol 28, Nr 12, pp 1403-1406 (USSR)
 ABSTRACT: This is a report on the plenary meetings held on the 8th Mendeleev Congress from March 16 to 23, 1959, in Moscow. From the lectures held on the plenary meetings, the following ones are mentioned: V. S. Fedorov, Chairman of the State Committee of the Council of Ministers of the USSR for Chemistry (Ref 3); V. A. Kargin (Ref 4); A. N. Nesmeyanov (Ref 5); N. N. Semenov (Ref 6); V. I. Spitsyn (Ref 7); A. P. Vinogradov (Ref 8); V. A. Engel'gardt (Ref 9); A. V. Sokolov (Ref 10); Director NIIKhimash V. B. Nikolayev (Ref 11); Ya. K. Syrkin (Ref 12) and A. P. Aleksandrov. All lectures were devoted to basic problems of modern chemistry and technology as well as to perspective problems of science. Plenary meetings were held with: Academician A. N. Nesmeyanov, Professor I. P. Losev, Academician A. Ye. Arbuzov, Academician S. I. Vol'fkovich, Corresponding Member AS USSR N. M. Zhavoronkov, Professor V. M. Kakabadze, Academician N. N. Semenov in the chair. The activity of the ✓

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8th Mendeleev Congress. Plenary Meetings

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sections took place from March 17 to 23, in the premises of Moscow University, at the Institut biologicheskoy i meditsinskoy khimii Akademii Meditsinskikh nauk SSSR (Institute of Biological and Medical Chemistry of the Academy of Sciences, USSR), and at the Nauchnyy institut udobreniy i insektovungitsidov im. Ya. V. Samoylova (Scientific Institute of Fertilizers and Insecto-fungicides imeni Ya. V. Samoylov) (Ref 13). ✓

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SOV/74-28-12-3/25

AUTHORS: Kozlov, V. V., Vol'fkovich, S. I.

TITLE: 8th Mendeleev Congress. Section of Inorganic Chemistry and Technology

PERIODICAL: Uspekhi khimii, 1959, Vol 28, Nr 12, pp 1406-1408 (USSR)

ABSTRACT: The Section of Inorganic Chemistry and Technology was working under the guidance of Academician I. I. Chernyayev, of the Deputies: Academician V. I. Spitsyn, Corresponding Member AS USSR B. V. Nekrasov, Scientific Secretary Candidate of Chemical Sciences O. N. Andrianova. Secretaries of the Section were: Candidate of Chemical Sciences T. N. Dymova, Candidate of Chemical Sciences I. D. Kolli, Candidate of Chemical Sciences A. I. Lazareva, Ye. I. Ionova, and V. T. Orlova. The meetings were held with: Academicians I. I. Chernyayev, A. A. Grinberg, and I. V. Tananayev; Corresponding Member AS USSR I. A. Kazarnovskiy, Corresponding Member AS USSR B. V. Nekrasov, Doctor of Chemical Sciences I. N. Lepeshkov, Corresponding Member AS USSR A. V. Novoselova, Professor G. A. Meyerson, Professor O. Ye. Zvyagintsev in the chair. Lectures were held by: V. I. Spitsyn, I. V. Yanitskiy, Ye. A. Ippolitova, I. A. Kazarnovskiy, A. V. Novoselova, K. N. Semenenko, A. I. Grigor'yev, S. Z. Ma-

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SOV/74-28-12-3/25

8th Mendeleev Congress. Section of Inorganic Chemistry and Technology

karov, L. G. Berg, V. I. Mikheyeva, N. N. Sirota, B. A. Berem-
zhanova, L. A. Borovskikh, Ya. Ye. Vil'nyanskiy, S. I. Vol'-
fkovich, N. N. Postnikov, A. A. Ionass, V. V. Illarionov, R. Ye.
Remen, Ye. P. Ozhigov, G. A. Meyerson, V. M. Lekaye, A. G. Ka-
satkin, L. N. Yelkin, Ye. Ya. Vil'nyanskiy, Ye. I. Savintsova,
L. A. Borovskikh, A. I. Teterovkov, L. S. Bychikhina, I. I.
Chernyayev, L. A. Nazarova, V. S. Orlova, A. A. Grinberg, V. G.
Tronev, K. B. Yatsimirskiy, B. V. Ptitsyn, D. I. Vinogradova,
Ye. N. Tekster, L. N. Sheronov, Z. A. Shek, Ye. Ye. Kriss,
O. I. Zakharov-Nartsissov, O. Ye. Zvyagintsev, V. I. Spitsyn,
Ye. A. Ippolitova, A. P. Sokolov, V. M. Vdovenko, Ye. A. Smir-
nova, D. N. Suglobov, L. M. Gindin, P. I. Bobikov, E. F. Koub,
I. F. Kopp, A. M. Rozen, N. P. Ter-Oganesov, N. I. Zagarskaya,
V. A. Kargin, R. P. Lastovskiy, T. A. Matveyeva, Yu. V. Shirokiy.
The lectures gave a survey of work carried out in the field of
inorganic chemistry. New methods developed for the investigation
of inorganic systems and new procedures designed for the intro-
duction into industry were reported.

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SOV/74-28-12-4/25

AUTHORS: Kozlov, V. V., Vol'fkovich, S. I.

TITLE: 8th Mendeleev Congress.. Section of Organic Chemistry and Technology

PERIODICAL: Uspekhi khimii, 1959, Vol 28, Nr 12, pp 1408-1415 (USSR)

ABSTRACT: The Section (Ref 15) was working under the guidance of Academician B. A. Kazanskiy, Deputies: Corresponding Member AS USSR P. A. Moshkin, Doctor of Chemical Sciences V. N. Belov, Scientific Secretary: Candidate of Chemical Sciences B. I. Stepanov. Activity of the Section took place, besides plenary meetings, within the following 4 subsections: 1) Catalytic Reactions and Some General Problems of Organic Chemistry and Technology (Head: Academician B. A. Kazanskiy, Scientific Secretary: I. V. Gostunskaya). The chairmen of this subsection were: Professor D. V. Sokolovskiy, Academician AS Ukrainskaya SSR Ye. A. Shilov, Corresponding Member Uzbekskaya SSR I. P. Tsukervanik, Academician A. A. Balandin, Academician AS Azerbaydzhan SSR Yu. G. Mamedaliyev, Corresponding Member AS USSR A. D. Petrov, Professor Yu. A. Gorin, Professor B. V. Tronov. 2) Chemistry and Technology of Aliphatic and Alicyclic Compounds (Head: ✓

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8th Mendeleev Congress. Section of Organic Chemistry and Technology

Corresponding Member AS USSR P. A. Moshkin, Scientific Secretary: L. N. Nokhapetyan). Chairmen: Corresponding Member AS USSR P. A. Moshkin, Professor V. N. Belov, Corresponding Member AS USSR S. N. Danilov, Corresponding Member AS USSR G. A. Razuvaev, Academician B. A. Arbuzov, Professor I. Ya. Postovskiy, Professor T. I. Temnikova, Academician Azerbaydzhan SSR Yu. G. Mamedaliyev. 3) Chemistry and Technology of Aromatic and Heterocyclic Compounds (Head: Professor Yu. K. Yur'yev, Scientific Secretary: I. I. Grandberg). Chairmen: Corresponding Member AS USSR N. N. Vorozhtsov, Professor B. A. Poray-Koshits, Professor A. A. Spryskov, Academician Ukrainskaya SSR A. I. Kiprianov, Professor V. A. Izmail'skiy, Professor Yu. K. Yur'yev, Professor V. V. Kozlov. 4) Chemistry and Technology of Elemental Organic Compounds (Head: Corresponding Member AS USSR R. Kh. Freydlina, Scientific Secretary: Ye. I. Vasil'yeva). Chairmen: Corresponding Member AS USSR R. Kh. Freydlina, Corresponding Member AS USSR O. A. Reutov, Academician M. I. Kabachnik, Professor A. A. Petrov, Professor G. Kh. Kamay, Corresponding Member AS USSR D. N. Kursanov, Corresponding Member AS USSR K. A. Kocheshkov. Lectures held by G. V. Uvarov, ✓

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8th Mendeleev Congress. Section of Organic Chemistry and Technology

Kh. K. Ingol'd, G. A. Razuvayev, M. I. Kabachnik, Ye. A. Shilov,
I. V. Smirnov-Zamkov, G. A. Piskovitina, V. G. Ostroverkhov,
D. F. Mironova, G. F. Dvorko, A. A. Akhrem, A. V. Kamernitskiy,
L. D. Bergel'son, L. P. Badenkov, Ye. I. Klabunovskiy, A. A.
Balandin, B. A. Kazanskiy, I. V. Gastunskaya, A. I. Leonova,
A. F. Plate, V. I. Stanko, A. A. Balandin, V. I. Spitsyn, N. P.
Dobrosel'skaya, I. Ye. Mikhaylenko, R. M. Flid, Yu. A. Gorin,
I. K. Gorn, A. A. Balandin, M. L. Khidekel', V. V. Patrikeyev,
Ye. G. Vol'pova, A. V. Lyuter, E. M. Koganova, A. A. Vvedenskiy,
T. Ye. Shakhova, A. Ye. Panitkova, A. R. Perel'man, I. M. Dol-
gopol'skiy, A. L. Klebanskiy, Z. A. Dobler, M. Ya. Rubanik,
S. V. Zavgorodney, T. B. Gonsovskaya, L. S. Shvetsova, V. I.
Sidel'nikova, V. G. Vakhtin, Ye. A. Vdovtsova, Yu. G. Mamedali-
yev, I. P. Tsukervanik, A. A. Balandin, I. R. Konenko, A. A.
Tolstopyatova, Ye. I. Karpeyskaya, A. P. Rudenko, A. A. Kuz'-
menko, Ya. L. Gol'dfarb, G. A. Rudakov, I. Ya. Postovskiy, N. N.
Vereshchagina, L. F. Trefilova, E. I. Chertkova, I. A. Korshu-
nov, N. F. Novotorov, N. A. Pestunovich, V. N. Dubovskaya, M. R.
Leonov, V. V. Perekalin, K. B. Rall', G. D. Padva, Yu. V. Svet-
kina, N. A. Dayev, V. M. Dashunin, R. Ya. Levina, V. R. Skvar-

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8th Mendeleev Congress. Section of Organic Chemistry and Technology

chenko, Yu. S. Shabarov, A. L. Klebanskiy, V. F. Vosik, Yu. A. Zhdanov, I. V. Machinskaya, V. A. Barkhash, R. Ya. Levina, V. K. Daukshas, P. A. Kaykaris, N. K. Kochetkov, L. I. Kudryashov, B. P. Gottikh, S. N. Danilov, V. F. Kazimirova, A. A. Lopatenok, T. I. Temnikova, B. A. Arbuzov, A. I. Konvalova, M. N. Shchukina, V. G. Yashunskiy, M. S. Malinovskiy, A. G. Yudasina, M. B. Blinova, B. L. Moldavskiy, V. G. Babel', R. I. Rudakova, M. Sh. Usmanova, V. K. Tsyskovskiy, P. A. Moshkin, R. I. Kobzova, V. V. Nesmelov, N. M. Labedeva, N. N. Terpilovskiy, O. V. Maminov, R. G. Danyushevskaya, N. I. Kobozev, Ye. N. Yeregin, M. F. Shostakovskiy, A. V. Bogdanova, N. A. Nikolayeva, G. Kh. Kamay, V. V. Perekalina, B. G. Yasnitskiy, Ye. B. Dol'berg, S. A. Sarkis'yants, Ts. I. Satanovskaya, A. P. Zaytsev, Sh. B. Aliyev, R. N. Degtyarenko, P. A. Moshkin, L. D. Pertsev, S. F. Kalinkin, A. A. Pryanishnikov, P. D. Borisov, A. N. Vodzinskaya, I. A. Grigorov, S. O. Skvortsov, V. P. Sumarokov, I. F. Chistov, S. V. Chepigo, M. Ye. Shpuntova, Ye. Ye. Shnayder, N. A. Vasyunina, G. S. Barysheva, Ye. S. Grigoryan, M. Z. Geras'kina, V. I. Isagulyants, L. N. Lavrishcheva, N. M. Przhiyalgovskaya, N. N. Vorozhtsov, A. A. Ponomarenko, V. A. Izmail'skiy, P. A. ✓

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8th Mendeleev Congress, Section of Organic Chemistry and Technology

Solodkov, N. A. Kitrasskiy, A. A. Spryskov, A. I. Kiprianov,
A. I. Tolmachev, N. S. Dokunikhin, L. M. Yegorova, G. S. Lisen-
kova, B. A. Poray-Koshits, V. V. Kozlov, B. I. Belov, B. I.
Stepanov, I. P. Gragerov, M. P. Ponomarchuk, M. M. Aleksankin,
A. F. Rekasheva, E. P. Dar'yeva, B. A. Geller, V. V. Perekalin,
T. A. Abramovich, I. P. Gragerov, B. I. Kissin, D. M. Ushakov,
P. K. Krutkov, Z. I. Krutikova, Ye. M. Chernysheva, G. M. Bar-
kov, N. K. Moshchinskaya, Yu. N. Sheynker, S. A. Giller, L. A.
Pavlova, E. D. Venus-Danilova, A. Fabritsy, M. N. Shchukina,
K. M. Murav'yeva, Yu. K. Yur'yev, K. Yu. Novitskiy, M. I. Far-
berov, B. F. Ustavshchikov, A. M. Kut'in,, A. N. Nesmeyanov,
O. A. Reutov, N. S. Kochetkova, N. A. Vol'kenau, V. D. Vil'-
chevskaya, V. G. Yashunskiy, V. F. Vasil'yeva, R. Kh. Freydlina,
A. B. Belyavskiy, A. A. Petrov, Kh. V. Bal'yan, Ye. I. Vasil'-
yeva, Sh. A. Karapetyan, N. A. Semenov, R. G. Petrova, V. N.
Kost, T. T. Sidorova, S. T. Ioffe, Yu. N. Sheynker, T. A.
Masteryukov, K. A. Kocheshkov, N. I. Sheverdina, T. V. Talalayeva,
Ye. M. Panov, L. V. Abramova, V. N. Setkina, D. N. Kursanov,
Ye. V. Lykova, I. P. Beletskaya, O. A. Reutov, R. Ye. Mardal-
eyshvili, E. M. Braynina, A. N. Nesmeyanov, O. V. Nogina, Yu. P.

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8th Mandeleev Congress. Section of Organic Chemistry and Technology

Kudryavtsev, L. I. Zakharkin, I. L. Knunyants, R. N. Sterlin, A. N. Nesmeyanov, L. S. Isayeva, T. P. Tolstaya, B. M. Mikhaylov, P. M. Aronovich, A. N. Blokhina, T. K. Kozminskaya, T. V. Kostroma, N. S. Fedotov, T. A. Shchegoleva, V. F. Velichko, T. P. Tolstaya, A. N. Nesmeyanov, L. S. Isayeva, L. G. Makarova, O. A. Ptitsyna, O. A. Reutov, A. E. Shipov, T. Ya. Medved', Ye. N. Tsvetkova, G. Kh. Kamay, V. S. Balabukh, G. K. Kozlova, Ye. A. Mironova, L. I. Tikhonova are mentioned. In the resolutions approved of by the Section, it was noted that lectures and communications give evidence of a noteworthy progress in scientific research work in the field of organic chemistry and technology which are being performed in the different republics and towns of the USSR. The Section calls the attention of the Management of the All-Union Chemical Society imeni D. I. Mendeleeva and the Office of the Department of Chemical Sciences of the AS USSR to the necessity of driving ahead work to establish a uniform nomenclature of organic compounds. It is necessary to publish, in addition, a series of monographs on methods of synthesis for elemental organic compounds as well as a specialized textbook designed for this field. ✓

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SOV/74-28-12-5/25

AUTHORS:

Kozlov, V. V., Vol'fkovich, S. I.

TITLE:

8th Mendeleev Congress. Section of Analytical Chemistry

PERIODICAL:

Uspekhi khimii, 1959, Vol 28, Nr 12, pp 1415-1417 (USSR)

ABSTRACT:

The Section of Analytical Chemistry (Ref 17) was working under the guidance of: Corresponding Member AS USSR I. P. Alimarin, Deputy: Doctor of Chemical Sciences A. I. Busev, Secretaries: Z. I. Podgayskaya, G. N. Bilimovich. Academician AS Ukrainskaya SSR A. K. Babko, Doctor of Chemical Sciences A. K. Ruzhentseva, Professor V. I. Petrashen', Doctor of Chemical Sciences V. I. Kuznetsov, Doctor of Chemical Sciences Yu. S. Lyalikov, Doctor of Physical and Mathematical Sciences A. K. Rusanov, Doctor of Chemical Sciences D. I. Ryabchikov, Corresponding Member of the Kazakhskaya SSR M. T. Kozlovskiy, Professor V. A. Nazarenko presided over the individual meetings of the Section. Lectures held by: I. P. Alimarin, I. V. Tananayev, K. B. Yatsimirskiy, Ye. F. Naryshkina, L. P. Rayzman, Yu. A. Klyachko, V. M. Peshkova, V. M. Bochkova, V. I. Kuznetsov, A. I. Busev, M. I. Ivanyutin, A. I. Portnov, R. P. Lastovskiy, Yu. I. Vaynshteyn, N. M. Dyatlova, V. Ya. Temkina, I. D. Kalpakova, V. A. Nazarenko, ✓

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SOV/74-26-12-5/25

8th Mendeleev Congress. Section of Analytical Chemistry

Ye. A. Biryuk, G. G. Shitareva, L. I. Kononenko, V. K. Kuznetsova, N. A. Tananayeva, V. A. Obolonchik, I. P. Alimarin, G. N. Bilimovich, D. I. Ryabchikov, A. N. Yermakov, V. K. Belyayeva, N. M. Marov, A. K. Babko, V. F. Toropova, Z. Kh. Shakhova, R. K. Motorkina, S. A. Gavrilova, Ye. N. Semenovskiy, A. I. Kokorin, N. A. Polotebnova, E. Ye. Vaynshteyn, Yu. I. Belyayev, V. V. Korolev, N. S. Poluektov, A. G. Karabash, Sh. I. Payzulayev, L. I. Moseyev, Yu. V. Morachevskiy, I. A. Stolyarova, M. O. Korshun, N. E. Gel'man, K. I. Glazova, N. S. Sheveleva, N. I. Larina, V. A. Klimova, Ye. G. Bereznitskaya, Ye. N. Merkulova, S. I. Sinyakova, Z. B. Rozhdestvenskaya, I. A. Yarovoy, Ya. P. Gokhshteyn, Yu. S. Lyalikov, M. B. Bardin, Yu. S. Temyanko, I. D. Panchenko, N. I. Udal'tsova, P. N. Paley, M. M. Senyavin, N. K. Galkina, A. M. Sorochan, N. G. Polyanskiy, A. S. Vernidub, V. I. Petrashen', A. A. Zhukhovitskiy, P. M. Turkel'taub are mentioned. The lectures dealt with chief trends in analytical chemistry: physical and physico-chemical analytical methods, use of new organic reagents, organic elementary microanalysis, chromatographic analysis, ✓

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SOV/74-28-12-5/25

8th Mendeleev Congress. Section of Analytical Chemistry

use of tagged atoms in analytical chemistry. ✓

Card 3/3

5(0)
 AUTHORS: Kozlov, V. V., Vol'fkovich, S. I. SOV/74-28-12-6/25
 TITLE: 8th Mendeleev Congress. Section of Physical Chemistry
 PERIODICAL: Uspekhi khimii, 1959, Vol 28, Nr 12, pp 1417-1419 (USSR)
 ABSTRACT: The Section of Physical Chemistry (Ref 18) was working under the guidance of: Academician V. N. Kondrat'yev, Deputy: Corresponding Member AS USSR V. V. Voyevodskiy, Corresponding Member AS USSR K. V. Chmutov, Corresponding Member AS USSR G. K. Boreskov, Scientific Secretary: Candidate of Chemical Sciences V. Ya. Shlyapintokh, Secretaries: Candidate of Chemical Sciences D. G. Knorre, Candidate of Chemical Sciences V. I. Vedeneyev. The individual meetings of the Section were held with Academician V. N. Kondrat'yev, Academician A. A. Balandin, Corresponding Member AS USSR V. V. Voyevodskiy, Corresponding Member AS USSR K. V. Chmutov, Professor A. A. Sokolov, Corresponding Member AS USSR N. M. Emanuel', Corresponding Member AS USSR G. K. Boreskov, Professor N. B. Neyman, Corresponding Member AS USSR Ya. I. Gerasimov, Professor D. N. Frank-Kamenetskiy, Professor K. P. Mishchenko, Corresponding Member AS USSR S. Z. Roginskiy in the chair. Ac-

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6th Mendeleev Congress. Section of Physical Chemistry

tivity of the Section took place in 3 subsections: Kinetics of Chemical Reactions; Structure of the Matter; Catalysis and Adsorption. Lectures held by: V. N. Kondrat'yev, V. V. Voyevodskiy, N. M. Emanuel', M. B. Neyman, K. K. Andreyev, N. N. Semenov, A. B. Nalbandyan, L. V. Karmilova, N. S. Yenikolopyan, N. V. Topchiyev, I. V. Patsevich, V. Ya. Shtern, A. P. Ballod, T. V. Fedorova, S. I. Molchanova, Ya. B. Zel'dovich, A. S. Sokolik, L. A. Lovachev, A. N. Terenin, F. I. Vilesov, M. V. Vol'kenshteyn, P. P. Shorygin, Z. S. Yegorova, I. V. Aleksandrov, N. D. Sokolov, S. A. Al'tshuller, K. A. Valiyev, A. I. Rivkind, B. M. Kozyrev, P. G. Tishkov, V. I. Avvakumov, L. A. Blyumenfel'd, S. Ya. Frenkel', S. Ye. Bresler, A. I. Kitaygorodskiy, Z. V. Zvonkova, V. V. Tarasov, V. V. Voyevodskiy, A. V. Storonkin, A. G. Morachevskiy, M. P. Susarev, M. M. Shul'ts, M. I. Usanovich, S. S. Urazovskiy, D. S. Tsiklis, Ya. I. Tur'yan, M. M. Dubinin, V. F. Kiselev, K. G. Krasil'nikov, A. V. Kiselev, A. Ya. Korolev, K. D. Shcherbakova, S. Z. Roginskiy, K. G. Boreskov, D. V. Sokol'skiy, Ya. B. Gorokhovatskiy, A. A. Balandin, V. E. Vasserberg, M. P. Maksimova, T. V. Georgiyevskaya, V. Kh. Matyushenko, A. A. Tolstopyatova,

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8th Mendeleev Congress. Section of Physical Chemistry

M. V. Polyakov, V. V. Shal', Z. Z. Vysotskiy, A. M. Rubinshteyn, M. T. Rusov, A. A. Slinkin, V. I. Yakerson are mentioned. In a resolution taken by the Section, it was noted that progress in theoretical and experimental investigations in the field of topochemical reactions is staying somewhat behind the general level of work in the field of chemical kinetics. Too little physical and physico-chemical methods are being used. It would be more appropriate to establish new laboratories and a central head office. ✓

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SOV/74-28-12-8/25

AUTHORS:

Kozlov, V. V., Vol'fkovich, S. I.

TITLE:

8th Mendeleev Congress. Section of Chemistry and Technology of Polymers

PERIODICAL:

Uspekhi khimii, 1959, Vol 28, Nr 12, pp 1422-1423 (USSR)

ABSTRACT:

The Section of Chemistry and Technology of Polymers¹ (Refs 21, 48-49,54) was working under the guidance of: Academician V. A. Kargin; Deputy: Professor Z. A. Rogovin; Scientific Secretary: Candidate of Chemical Sciences Yu. M. Malinskiy. Academician V. A. Kargin, Academician S. S. Medvedev, Professor Z. A. Rogovin, Professor N. N. Shorygina, Professor G. L. Slonimskiy, Corresponding Member AS USSR V. V. Korschak presided over the individual meetings of the Section. Lectures held by: I. P. Losev, L. A. Datskevich, K. D. Petrov, O. K. Gosteva, V. I. Pukhova, K. A. Andrianov, A. A. Zhdanov, O. Ya. Fedotova, N. I. Skripchenko, I. P. Losev, G. I. Kudryavtsev, Ye. A. Vasil'yeva-Sokolova, M. A. Zharkova, S. N. Ushakov, V. A. Kargin, P. V. Kozlov, N. A. Plate, I. I. Konoreva, Ye. V. Kuznetsov, S. S. Spasskiy, A. V. Tokarev, M. A. Mikhaylov, A. I. Tarasov, T. V. Molchanov, M. Ye. Mat'kov, V. V. Korshak, S. L. Sosin, M. V.

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SOV/74-28-12-8/25

8th Mendeleev Congress. Section of Chemistry and Technology of Polymers

Chistyakova, M. S. Akutin, B. M. Kovarskaya, L. I. Golubenkova,
K. N. Vlasova, S. P. Kalinina, N. Ya. Parlashkevich, V. N.
Kotrelev, Z. N. Tarasova, M. Ya. Kaplunov, N. A. Klauzen, B. A.
Dogadkin, V. L. Karpov, V. K. Lyubeznikov, M. G. Zaripova,
Yu. L. Margolina, B. A. Dogadkin, O. N. Belyatskaya, M. S.
Fel'dshteyn, I. I. Eytingon, Z. N. Nudel'man, A. S. Novikov,
K. F. Kaluzhenina, N. P. Zinchenko, G. N. Buyko, N. M. Arenzon,
A. I. Tumanova, V. A. Kargin, A. S. Novikov, F. A. Galii-Ogly,
G. M. Bartenev, N. V. Zakharenko, F. S. Tolstukhina, A. S. Kuz'-
minskiy, T. G. Degteva, A. A. Vansheydt, N. N. Kuznetsova, F. T.
Shostak, Z. A. Rogovin, V. A. Derevitskaya, N. V. Mikhaylov,
Z. G. Serebryakova, N. N. Shorygina, A. A. Chuksanova, A. F.
Semechkina, L. L. Sergeyeva, A. B. Pakshver, L. V. Kozlov,
V. I. Selikhova, G. S. Markova, V. A. Kargin, V. A. Kargin, V. A.
Kabanov, I. Yu. Marchenko, V. Ye. Gul' are mentioned. ✓

Card 2/2

5(0)
AUTHORS: Kozlov, V. V., Vol'fkovich, S. I. SOV/74-28-12-11/25

TITLE: 8th Mendeleev Congress. Section of Chemistry and Chemical Technology of Fuels

PERIODICAL: Uspekhi khimii, 1959, Vol 28, Nr 12, pp 1426-1428 (USSR)


ABSTRACT: The Section of Chemistry and Chemical Technology of Fuels (Ref 24) was working under the guidance of: Corresponding Member AS USSR N. I. Shuykin; Deputy: Corresponding Member AS USSR N. M. Karavayev; Scientific Secretary: M. A. Ryashentseva. The activity of the Section was divided into two topics: Petroleum Chemistry, and Coal Chemistry. Corresponding Member AS USSR N. I. Shuykin, Corresponding Member AS USSR N. M. Karavayev, Doctor of Chemical Sciences N. V. Lavrov, Doctor of Chemical Sciences N. G. Titov, Corresponding Member AS USSR A. D. Petrov, Candidate of Chemical Sciences S. I. Khromov, Professor A. F. Plate, Doctor of Chemical Sciences S. R. Sergiyenko, Professor N. I. Chernozhukov, Candidate of Technical Sciences V. P. Sukhanov presided over the individual meetings of the Section. Lectures held by: V. P. Sukhanov, N. M. Karavayev, N. V. Lavrov, N. I. Shuykin, N. G. Bekauri, A. F. ✓

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SOV/74-28-12-11/25

8th Mendeleev Congress. Section of Chemistry and Chemical Technology of Fuels

Plate, S. R. Sergiyenko, A. D. Petrov, V. I. Isagulyants, G. N. Maslyanskiy, V. A. Kobelev, N. R. Bursian, M. I. Ryskin, A. S. Fomina, L. Ya. Pobul', Z. A. Degtyareva, A. A. Kruglikov, M. A. Menkovskiy, A. N. Aleksandrova, K. Yu. Volkov, S. A. Gordon, L. V. Petrovskaya, V. M. Ratynskiy, T. I. Sendul'skaya, N. I. Shuykin, T. I. Naryshkina, I. M. Artyukhov, D. I. Zul'fugarly, B. A. Kazanskiy, S. I. Khromov, Ye. S. Balenkova, N. A. Seidova, Al. A. Petrov, S. V. Adel'son, T. Kh. Melik-Akhnazarov, I. I. Mukhin, D. I. Orochko, N. A. Chepurov, V. N. Kozlov are mentioned.



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SOV/74-28-12-13/25

AUTHORS:

Kozlov, V. V., Vol'fkovich, S. I.

TITLE:

8th Mendeleyev Congress. Section of Chemistry and Technology of Silicates

PERIODICAL:

Uspekhi khimii, 1959, Vol 28, Nr 12, pp 1429-1431 (USSR)

ABSTRACT:

The Section of Chemistry and Technology of Silicates (Refs 26, 57) was working under the guidance of: Academician AS UkrSSR P. P. Budnikov; Deputies: Academician of the Akademiya stroitel'stva i arkhitektury (Academy of Civil Engineering and Architecture) N. A. Toropov, Corresponding Member AS Ukrainskaya SSR O. P. Mchedlov-Petrosyan; Scientific Secretary: Candidate of Technical Sciences I. V. Kravchenko. Secretaries were: A. Ye. Rempel', N. V. Petrovykh, V. V. Myshlyayeva. Academician AS Ukrainskaya SSR P. P. Budnikov, Academician of the Academy of Civil Engineering and Architecture N. A. Toropov, Professor Yu. M. Butt, Professor V. V. Stol'nikov, Professor I. I. Kitaygorodskiy, Professor I. F. Ponomarev, Docent S. M. Royak presided over the individual meetings. The following topics were treated in the lectures: General Silicate Chemistry; Chemistry and Technology of Ceramics and Refractories; Binders and Glass.

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SOV/74-28-12-13/25

8th Mendeleev Congress. Section of Chemistry and Technology of Silicates

Lectures held by: P. P. Budnikov, Yu. M. Butt, S. K. Dubrovo, I. I. Kitaygorodskiy, Yu. A. Shmidt, V. A. Ioffe, M. K. Gal'perin, V. I. Minenko, S. M. Petrova, P. P. Budnikov, V. G. Savel'yev, O. M. Astreyeva, V. I. Guseva, N. S. Popov, N. S. Zavgorodniy, K. S. Kutateladze, N. G. Dzhincharadze, I. V. Kravchenko, Ye. V. Podushko, L. N. Rashkevich, Yu. M. Butt, V. F. Fedoryakin are mentioned. It is noted in the resolution that the most important tasks in the field of chemistry and technology for the next years are the extension of research work and the improvement of operational processes. ✓

Card 2/2

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SOV/74-28-12-14/25

AUTHORS:

Kozlov, V. V., Vol'fkovich, S. I.

TITLE:

8th Mendeleev Congress. Section of Radiochemistry and Isotope Chemistry

PERIODICAL:

Uspekhi khimii, 1959, Vol 28, Nr 12, pp 1431-1433 (USSR)

ABSTRACT:

The Section of Radiochemistry and Isotope Chemistry (Ref 27) was working under the guidance of Academician A. P. Vinogradov; Deputy: Academician AS Ukrainskaya SSR A. I. Brodskiy; Scientific Secretaries: Candidate of Chemical Sciences A. N. Yermakov, Candidate of Chemical Sciences G. A. Nekrasova. Academician A. P. Vinogradov, Corresponding Member AS USSR I. Ye. Starik, Academician V. I. Spitsyn, Academician S. S. Medvedev, Academician AS Ukrainskaya SSR A. I. Brodskiy, Corresponding Member AS USSR N. M. Zhavoronkov presided over the meetings of the Section. Lectures held by: A. P. Vinogradov, A. K. Lavrukhnina, S. S. Rodin, A. A. Pozdnyakov, I. P. Alimarin, Yu. A. Zolotov, Yu. P. Novikov, P. N. Paley, Ye. S. Pal'shin, G. N. Yakovlev, I. P. Semenov, G. N. Flerov, A. D. Gel'man, Yu. B. Gerlit, V. M. Vdovenko, M. P. Koval'skaya, T. V. Kovaleva, A. A. Lipovskiy, M. G. Kuzina, L. N. Lazarev, S. Yu. Yelovich, B. K. ✓

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SOV/74-28-12-14/25

8th Mendeleev Congress. Section of Radiochemistry and Isotope Chemistry

Preobrazhenskiy, A. V. Kalyamin, O. M. Lilova, I. Ye. Starik,
V. I. Grebenshchikova, R. V. Bryzgalova, M. B. Chernyavskaya,
V. I. Bobrova, K. B. Zaborenko, I. V. Kolosov, V. V. Fomin,
An. N. Nesmeyanov, Ye. A. Borisov, I. Zvara, B. G. Dzantiyev,
B. A. Kuznetsov, A. D. Maliyevskiy, Ts. I. Zal'kind, N. B.
Miller, G. S. Tyurikov, G. Z. Gochaliyev, V. I. Veselovskiy,
N. A. Bakh, L. G. Bugayenko, V. I. Medvedovskiy, A. A. Revina,
L. S. Polak, A. V. Topchiyev, N. Ya. Chernyak, S. Ya. Pshe-
zhetskiy, M. A. Proskurnin, Ye. A. Shilov, F. M. Vaynshteyn,
L. L. Strizhak, A. I. Brodskiy, K. I. Sakodyskiy, S. I. Bobkov,
N. M. Zhavoronkov, A. A. Balandin,, V. I. Spitsyn, L. I. Bar-
kova, V. I. Duzhenkov, Yu. S. Lazurkin, M. A. Makul'skiy, N. N.
Sevryugova, O. V. Uvarov, S. I. Babkov, G. N. Chernykh, V. A.
Sokol'skiy, N. M. Zhavoronkov, I. B. Rabinovich, N. N. Tun-
tskiy, M. V. Gur'yev, M. V. Tikhomirov, V. L. Tal'roze, Ye. L.
Frankovich are mentioned. Out-of-program lectures were held by:
I. P. Selinov, N. N. Krot, A. G. Kozlov, V. P. Shvedov, A. V.
Stepanov, M. M. Senyavin, I. Ya. Petrov, V. L. Karpov, V. A.
Nikishina, V. P. Meshcheryakov, B. S. Kir'yanov, A. P. Smirnov-
Averin, B. V. Ershler, M. A. Nezhevenko, G. G. Misishcheva,
I. V. Vereshchinskiy, A. K. Pikayev, P. Ya. Glazunov, Ya. M.

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SOV/74-28-12-14/25

8th Mendeleev Congress. Section of Radiochemistry and Isotope Chemistry

Varshavskiy, V. L. Karpov, I. Ya. Petrov, Z. S. Bulanovskaya,
A. M. Rozen, A. B. Anufriyev, A. D. Bondar', N. G. Zaytseva,
Lo Weng-chung. ✓

Card 3/3

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SCV/74-28-12-15/25

AUTHORS: Kozlov, V. V., Vol'fkovich, S. I.

TITLE: 8th Mendeleev Congress. Section of Theoretical and Applied Electrochemistry

PERIODICAL: Uspekhi khimii, 1959, Vol 28, Nr 12, pp 1433-1436 (USSR)

ABSTRACT: The Section of Theoretical and Applied Electrochemistry (Ref 28) was working under the guidance of: Academician A. N. Frumkin; Deputy: Professor S. I. Sklyarenko; Scientific Secretary: Candidate of Chemical Sciences T. V. Kalish. Academician A. N. Frumkin, Professor Ya. M. Kolotyrkin, Corresponding Member AS USSR and President of the Akademiya nauk Litovskoy SSR (Academy of Sciences of the Lithuanian SSR) Yu. Yu. Matulis, V. I. Veselovskiy, Academician AS Ukrainskaya SSR Yu. K. Delimarskiy, Professor Ya. V. Durbin, Doctor of Technical Sciences L. M. Yakimenko, Professor N. D. Tomashov, Professor B. N. Kabanov, Professor Yu. V. Baymakov presided over the individual meetings of the Section. Lectures held by: A. N. Frumkin, L. M. Yakimenko, Yu. V. Baymakov, S. I. Sklyarenko, V. V. Stender, R. I. Agladze, A. I. Nizhnik, N. A. Izmaylov, N. P. Nikol'skiy, M. M. Shul'ts, N. V. Poshekhonova, A. I. Parfenova, O. V. Mazurin, N. Ye. Khomutov, I. S. Golinker, V. G. Levich, Ya. M.

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SOV/74-28-12-15/25

8th Mendeleev Congress. Section of Theoretical and Applied Electrochemistry

Kolotyarkin, L. A. Medvedeva, N. V. Nikolayeva-Fedorovich, O. A. Petriy, A. N. Frumkin, S. V. Gorbachev, M. A. Loshkarev, I. P. Chernobayev, B. I. Tomilov, A. G. Stromberg, D. P. Zosimovich, B. N. Kabanov, A. T. Vagramyan, A. P. Popkov, N. T. Kudryavtsev, G. K. Smolenskaya, V. M. Karatayeva, R. G. Golovchanskaya, V. V. Andreyev, T. P. Stepanova, Ye. S. Volkova, I. V. Krotov, V. M. Berenblit, I. Ya. Lantratova, A. I. Shultin, G. I. Volkov, D. P. Semchenko, K. G. Il'in, M. F. Lantratov, A. Ya. Alabyshev, Yu. Yu. Matulis, K. M. Gorbunova, Yu. M. Polukarov, O. S. Popova, A. A. Sutyagin, G. M. Kamar'yan, I. L. Rozenfel'd, V. P. Maksimchuk, L. K. Lepin', A. Ya. Vayvade, A. K. Lokenbach, Yu. K. Delimarskiy, K. M. Kalabina, V. V. Kuz'movich, T. P. Kaptsova, S. I. Sklyarenko, V. I. Konstantinov, Ye. A. Markina, V. M. Amosov are mentioned. In the resolution, it is stressed that a broad-scale introduction of electrochemical production methods into industry is required, if the tasks of the 7 Years Plan are to be fulfilled. ✓

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SOV/74-28-12-16/25

AUTHORS:

Kozlov, V. V., Vol'fkovich, S. I.

TITLE:

8th Mendeleev Congress. Section of the Chemistry of Metals and Alloys

PERIODICAL:

Uspekhi khimii, 1959, Vol 28, Nr 12, pp 1436-1437 (USSR)

ABSTRACT:

The Section of Chemistry of Metals and Alloys (Ref 29) was working under the guidance of Corresponding Member AS USSR N. V. Ageyev; Deputy: Professor I. I. Kornilov; Scientific Secretary: Candidate of Technical Sciences R. S. Polyakova. Academician I. P. Bardin, Professor Ye. Ye. Cherkashin, Professor Ye. M. Savitskiy, Professor Ye. S. Makarov, Professor I. I. Kornilov, Corresponding Member AS USSR N. V. Ageyev presided over the individual meetings of the Section. Lectures held by: N. V. Ageyev, I. I. Kornilov, Ye. M. Savitskiy, A. M. Yevseyev, Ya. I. Gerasimov, A. V. Nikol'skaya, A. Ya. Shonyayev, K. G. Khomyakov, V. A. Troshkina, Yu. D. Tret'yakov, Ye. Ye. Cheburkova, T. A. Borscheva, M. F. Alekseyenko, Ye. L. Bushmanova, L. V. Zaslavskaya, S. T. Kishkin, G. N. Orekhov, A. F. Platonov, N. M. Popova, G. M. Rovenskiy, K. P. Sorokina, N. I. Blok, N. F. Lashko, M. N. Kozlova, Ye. S. Makarov, Ye. I. ✓

Card 1/2

SOV/74-28-12-16/25
8th Mendeleev Congress. Section of the Chemistry of Metals and Alloys

Gladyshevskiy, P. I. Kripyakevich, Yu. B. Kuz'ma, Yu. N. Andreyev, G. S. Zhdanov, V. K. Grigorovich are mentioned. A joint meeting together with the Section of Inorganic Chemistry and the Section of Physical Chemistry was held on March 18. Lecturers were: N. N. Sirota, B. F. Ormont, L. G. Dudkin, Z. G. Pinsker, Ya. A. Ugay, N. A. Goryunova. In the resolution, the necessity to extend considerably work in the field of physico-chemical analysis, the study of phase diagrams, and the crystal chemistry of metal alloys was stressed. ✓

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5(0)
AUTHORS: Kozlov, V. V., Vol'fkovich, S. I. SOV/74-28-12-7/25

TITLE: 8th Mendeleev Congress. Section of Colloid Chemistry

PERIODICAL: Uspekhi khimii, 1959, Vol 28, Nr 12, pp 1419-1422 (USSR)

ABSTRACT: The Section of Colloid Chemistry (Ref 19) was working under the guidance of: Academician P. A. Rebinder; Deputy: Doctor of Technical Sciences N. V. Mikhaylov; Scientific Secretary: Professor G. I. Fuks. Academician P. A. Rebinder, Academician AS Belorusskaya SSR N. F. Yermolenko, Professor Ye. M. Aleksandrova, Corresponding Member AS USSR B. V. Deryagin, Professor M. Ye. Shishniashvili presided over the individual meetings of the Section. Lectures held by: V. P. Smilga, B. V. Deryagin, N. A. Krotova, L. P. Morozova, Yu. F. Deynega, A. V. Dumanskiy, G. V. Vinogradov, I. Ye. Neymark, V. V. Karasev, B. V. Deryagin, G. I. Fuks, M. S. Ostrikov, S. I. Popel', N. V. Pertsov, N. F. Yermolenko, Z. A. Krivchik, A. B. Taubman, S. A. Nikitina, P. I. Yermilov, I. Ye. Neymark, P. V. Moryganov, B. N. Mel'nikov, V. F. Androsov, A. A. Kharkharov, V. F. Boyko, Ye. M. Aleksandrova, V. N. Tsvetkov, N. S. Razumikhina, A. I. Yurzhenko, R. V. Kucher, A. B. Taubman, A. F. Koretskiy, M. A. Kovbuz, S. M. /

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SOV/74-28-12-7/25

8th Mendel'ev Congress. Section of Colloid Chemistry

Levi, Z. Ya. Berestneva, V. A. Kargin, B. V. Deryagin, Yu. M. Glazman, D. N. Strazhesko, E. M. Natanson, S. Ye. Kharin, K. S. Lyalikov, P. M. Silin, P. A. Rebinder, Ye. Ye. Segalova, M. P. Volarovich, I. V. Churayev, N. V. Mikhaylov, Ye. Ye. Kalmykova, O. P. Mchedlov-Petrosyan, F. A. Latyshev, A. G. Bunakov, N. A. Levchuk, V. I. Yakimova, S. L. Talmud, K. P. Mishchenko, A. A. Morozov, S. N. Stavrov, L. I. Belen'kiy, M. Ye. Kazanskaya, T. V. Bromberg, P. A. Demchenko are mentioned. In the resolution taken by the Section, it was pointed to great progress achieved in applied theoretical colloid chemistry in the USSR. The attention of the OKhN AS USSR was called to the advantages of establishing an Institut fiziko-khimicheskoy mekhaniki (Institute of Physico-chemical Mechanics). ✓

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SOV/74-28-12-17/25

AUTHORS: Kozlov, V. V., Vol'fkovich, S. I.

TITLE: 8th Mendeleev Congress. Section of Economics, Planning, and Organization of Chemical Plants

PERIODICAL: Uspekhi khimii, 1959, Vol 28, Nr 12, pp 1437-1439 (USSR)

ABSTRACT: The Section of Economics, Planning, and Organization of Chemical Plants (Refs 30,59) was working under the guidance of: Corresponding Member AS USSR N. N. Nekrasov; Deputy: Professor N. P. Fedorenko; Scientific Secretary: A. L. Rabkina. Corresponding Member AS USSR N. N. Nekrasov, Professor S. A. Pervushin, Professor N. P. Fedorenko, Professor L. I. Raytburg, Professor A. I. Dzhen-Litovskiy presided over the individual meetings of the Section. Lectures held by: N. N. Nekrasov, P. A. Khokhryakov, G. F. Borisovich, A. I. Ravdel', N. A. Zeligman, V. G. Fridenberg, S. A. Pervushin, S. Ya. Rachkovskiy, N. P. Fedorenko, Ye. P. Shchukin, N. P. Fedorenko, G. F. Borisovich, P. A. Boris, A. L. Rabkina, I. V. Rostanin, G. Ye. Birger, Ye. P. Ivanova are mentioned. In its resolution, the Section has pointed to the fact that the fulfillment of the 7 Years Plan requires a detailed study of methods to be used for the most

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SOV/74-28-12-17/25

8th Mendeleev Congress. Section of Economics, Planning, and Organization of Chemical Plants

effective utilization of material and manpower resources. The attention of the Gosplan SSSR (State Planning Board of the USSR), Otdeleniye ekonomicheskikh nauk AN SSSR (Department of Economical Sciences of the AS USSR), Goskomitet Soveta Ministrov SSSR (State Committee of the Council of Ministers of the USSR), and Ministerstvo vysshego obrazovaniya SSSR (Ministry of Higher Education of the USSR) was called to the necessity to improve and to extend the treatment of economical problems in chemical industry. Moreover, the tasks of the Institutes of the Academy of Sciences and of the State Planning Office of the USSR were discussed. It was pointed to the necessity to promote training in economics at chemical colleges, and to increase the publication of specialized literature treating questions of economics and planning of chemical production. ✓

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5 (3)

AUTHOR:

Kozlov, V. V.

SOV/79-29-4-65/77

TITLE:

Investigations in the Anthraquinone Series (Issledovaniya v ryadu antrakhinona). XXVII. On the Hydrolysis of the 1,5- and 1,8-Nitrosulfo Acids of Anthraquinone (1,5- and 1,8-Nitroxyanthraquinone) (XXVII.0 gidrolize 1,5-i 1,8-nitrosul'fokislot antrakhinona (1,5-i 1,8-nitrooksiantrakhinony)).

PERIODICAL:

Zhurnal obshechey khimii, 1959, Vol 29, Nr 4, pp 1344 - 1349 (USSR)

ABSTRACT:

The authors continued their previous investigations of hydrolysis in the series of anthraquinone (Refs 1-4) and investigated in the present paper the influence exercised by the nitro group upon the hydrolysis of the sulfo group in anthraquinone, proceeding from the anthraquinone—nitrosulfo acids. Except a patent (Ref 5) nothing has hitherto been published on the hydrolysis of these compounds. This patent describes the hydrolysis of the 1-nitro-8-sulfo acid of anthraquinone in sulfuric acid of 50% with mercury at 190 - 200°, under formation of the α-nitroanthraquinone. It was found that the hydrolysis of the 1,5- and 1,8-nitrosulfo acids of anthraquinone in sulfuric acid and water without participation of mercury salts

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Investigations in the Anthraquinone Series . .

SOV/79-29-4-65/77

XXVII . On the Hydrolysis of the 1,5- and 1,8-Nitrosulfo
Acids of Anthraquinone (1,5- and 1,8-Nitroxanthraquinone)

is possible in the open as well as in the closed system at high temperature. The hydrolysis in water, in fused tubes, leads to the formation of the hitherto unknown 1,5- and 1,8-nitrosulfo acids of anthraquinone by the substitution of the sulfo group by hydroxyl in the initial product. The hydrolysis in concentrated sulfuric acid proceeds in the open system under formation of the α -nitroanthraquinone by the substitution of the sulfo group of the initial product by hydrogen. The hydrolysis with sulfuric acid in fused tubes proceeds under formation of a mixture of products which are produced by the substitution of the sulfo- as well as of the nitro group of the initial product by hydrogen and hydroxyl. The hydrolysis with concentrated sulfuric acid is in these two systems complicated by the oxidation processes. The presence of the anthraquinone-1-sulfo acid in the position 5, and especially in the position 8, inhibits considerably the mobility of the sulfo group, compared with the behaviour of the sulfo group in the 1-sulfo acid or the 1,5- and 1,8-disulfo acid of anthraquinone. The tables show the hydrolysis results under various reaction conditions (time, system,

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Investigations in the Anthraquinone Series. . SOV/79-29-4-65/77
XXVII) On the Hydrolysis of the 1,5- and 1,8-Nitrosulfo
Acids of Anthraquinone (1,5- and 1,8-Nitroxanthraquinone)

temperature, water volume, acid concentration. There are 5
tables and 8 Soviet references.

ASSOCIATION: Moskovskiy institut narodnogo khozyaystva imeni G. V. Plekhanova
(Moscow Institute of National Economy imeni G. V. Plekhanov)

SUBMITTED: February 8, 1958

Card 3/3

BROZOVSKIY, D.I.; KOZLOV, V.V.

Using trihydroxynaphthalenes as oxidation dyes for furs. Kozh.-
obuv.prom. 2, no.6:18-20 Je '60. (MIRA 13:9)
(Dyes and dyeing--Fur) (Naphthalenetriol)

LEBEDEV, N.N.; BALTADZHI, I.I.; KOZLOV, V.

Effect of the activity of catalysts on the relative reactivity
of toluene and benzene during chlorination. Zhur. VKHO 5
no. 2:236-237 '60. (MIRA 14:2)

1. Moskovskiy khimiko-tekhnologicheskii institut imeni
D.I. Mendeleeva.
(Toluene) (Benzene) (Chlorination)

KOZLOV, V.V.; BROZOVSKIY, D.I.; GAVRILOVA, V.M.

Naphthalene series. Part 19: 1,3,7- Trihydroxynaphthalene. Zhur.
ob.khim. 30 no.8:2714-2718 Ag '60. (MIRA 13:8)
(Naphthalene)

KOZLOV, V.V.; SILAYEVA, Ye.A.

Anthraquinone series. Part 31: Sulfonic acids of trans-dibenzopyrenequinone. Zhur. ob. khim. 30 no.11:3766-3772
N'60. (MIRA 13:11)

1. Moskovskiy institut narodnogo khozyaystva imeni G.V.Plekhanova.
(Dibenzopyrenequinone)

KOZLOV, V.V.; VESELOVSKAYA, I.K.

Naphthalene series. Part 20: Kinetics of the reaction of α -naphthol with aqueous solutions of ammonia and ammonium sulfite. Zhur. ob. khim. 30 no.12:4088-4092 D '60. (MIRA 13:12)

1. Moskovskiy institut narodnogo khozyaystva imeni G.V.Flekhanova.
(Naphthol) (Ammonia) (Ammonium sulfite)

KHRUSTALEVA, V.N.; KOZLOV, V.V.

Colorimetric determination of inulin. Izv. vys. ucheb. zav.;
pishch. tekhn. no.4:152-156 '61. (MIRA 14:8)

1. Moskovskiy institut narodnogo khozyaystva imeni G.V.Flekhanova,
laboratoriya organicheskoy khimii.
(Inulin) (Colorimetry)